

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: hydroxide is more basic than hydroxide in acid



Four different types of bonds

1. Ionic Bonds
2. Covalent Bonds
3. Hydrogen Bonds
4. Metallic Bonds

There are four types of bonding: ionic bonding and covalent bonding

- Ionic bonds are formed by the transfer of electrons from one atom to another (anions and cations) (opposite charges)



- Covalent bonds are formed by sharing of electrons between two atoms



The Polarity of the bond



Hydrogen and Van der Waals



Metals



There are two main categories of bonds: ionic and covalent

Definition of ionic bond



Polar and Nonpolar Covalent Bonds

Depends on the difference in the electronegativity



Nonpolar covalent bond



There are two types of covalent bonding: ionic and covalent bonding

- Ionic bonds are formed by the transfer of electrons from one atom to another (anions and cations) (opposite charges)



- Covalent bonds are formed by sharing of electrons between two atoms



Polar and non-polar covalent bonds

Covalent bonds are classified as polar and non-polar.

- Polar bonds are formed by unequal sharing of electrons.
- Non-polar bonds are formed by equal sharing of electrons.
- Very simple bond is a covalent bond.
- Non-polar bonds consist of two equal and identical atoms.



Polar bonds are stronger than non-polar

Polar and Nonpolar Covalent Bonds



Depends on the difference in the electronegativity



- When the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond (like a covalent bond between chlorine atoms for the Cl-Cl and C-Cl bonds).

- If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond (nonpolar covalent bond) electrons are shared by one atom (O, N, F and other halogens).

- A difference of 1.7 or higher one or both the atoms have a larger shared and a weaker bond is formed.

A Summary of Ring-opening Reactions of Epoxydes with Strong and Weak Nucleophiles



Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).

1. Nonpolar covalent bond
2. Polar covalent bond
3. Ionic bond



- The dipole moment is shown as a vector pointing from the less electronegative atom towards the more electronegative atom.
- The vector is longer the more partial negative charge and the less electronegative atom has a partial positive charge (δ+).
- The dipole moment indicates the difference in electronegativity.
- The dipole moment indicates the difference in electronegativity.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of the bond dipole moments. If the dipole moments are canceled, it just means a neutral dipole.



Sum of vector of the bond dipole moment for the molecule

Bond Length and Bond Strength

Reactions of bond length



Length: H-H > H-Cl > H-Br > H-I

Strength: H-H < H-Cl < H-Br < H-I

Ionic bond: $\text{Na}^+ \text{Cl}^-$ > $\text{H}^+ \text{Cl}^-$

Nonpolar covalent bond: H_2 > Cl_2

Polar covalent bond: H_2O > Cl_2

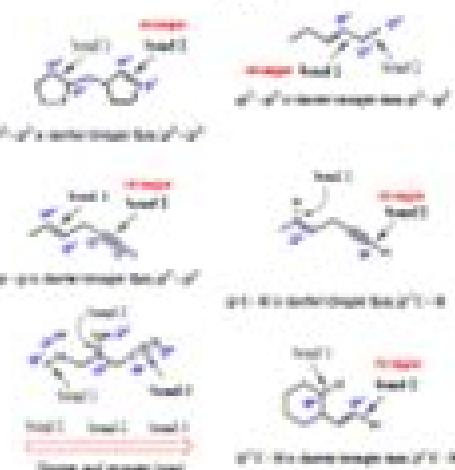
Reactions of bond strength

Average values of C-C and C-H bond depending on the molecule

C-C Bond	$\text{H}_2\text{C}=\text{CH}_2$	$\text{H}_2\text{C}=\text{CHCO}_2\text{C}_2\text{H}_5$	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$
Length	1.37 Å	1.37 Å	1.37 Å
Strength	112 kJ/mol	112 kJ/mol	107 kJ/mol
Hybridization	sp^2	sp^2	sp^3
Electronegativity	2.55	2.55	2.55
Length	1.10 Å	1.10 Å	1.10 Å
Strength	112 kJ/mol	112 kJ/mol	107 kJ/mol

Higher the bond length and stronger bond

Example



Properties of covalent bonds

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Arthur James Wells

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