

# Reactions of Amines and Epoxydes

## Preparation of Amines

Ammonium hydroxide:  $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



## Four different types of bonds



There are four types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electrons from one atom to another (anions and cations) (opposite charges).



## Covalent ionic and polar covalent bonds

- Covalent bonds are classified as ionic (polar) and non-polar.
- Ionic bonds are formed by the transfer of electrons.
- Non-polar bonds are formed by the sharing of electrons.
- Polar covalent bonds are formed by the sharing of electrons.
- Non-polar covalent bonds are formed by the sharing of electrons.



Non-polar covalent bonds are stronger than ionic.

## Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, polar bonds can be polar and nonpolar.

- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond (like a bond between chlorine and the C-C bond).
- If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond (nonpolar covalent bond) electrons are shared in a region of N, O, F and other halogens.
- A difference of 1.7 or higher one or both of the atoms will have shared and unequal shared in a polar bond is formed.

## A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Protagonists

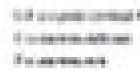


### Protagonists and Reactions

## Bond Types, Strength, Length, Polarity and Dipole Moment

### Dipole Moment

The polarity of a bond is denoted by the dipole moment ( $\mu$ ).



Directional dipole moment of the covalent bond.

- The dipole moment is shown as a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector is proportional to the partial negative charge and the partial positive charge.
- The dipole moment is zero for a polar covalent bond and for a nonpolar covalent bond.
- The dipole moment is zero for ionic bonds.

### Molecular Dipole

To measure the molecular polarity, the total sum of the dipole is used.

The polarity of the molecule is equal to the sum of the individual bond dipole moments. If the dipole moment is canceled, it just means a neutral dipole.



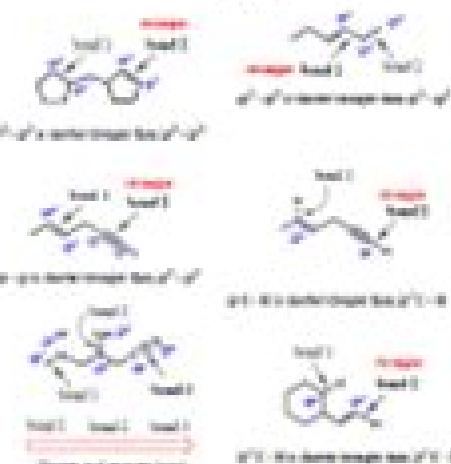
Sum of bond dipole moment of the molecule.

### Average values of C-C and C-H bond depending on the substituents

C-C Bond	$\text{H}_3\text{C}-\text{CH}_3$	$\text{H}_2\text{C}=\text{CH}_2$	$\text{CH}_3\text{CH}_2\text{CH}_3$
Hybridization	$\text{sp}^3$ hybrid	$\text{sp}^2$ hybrid	$\text{sp}^3$ hybrid
Length	1.54 Å	1.35 Å	1.53 Å
Strength	10.5 kJ/mol	17.1 kJ/mol	10.5 kJ/mol
Hybridization	$\text{sp}^3$	$\text{sp}^2$	$\text{sp}^3$
Strength	10.5 kJ/mol	17.1 kJ/mol	10.5 kJ/mol
Length	1.54 Å	1.35 Å	1.53 Å
Strength	10.5 kJ/mol	17.1 kJ/mol	10.5 kJ/mol

### Dipole moment and bond length

## Example



### Bond Length and Bond Strength

#### Relative bond length



Bond length:  $\text{C-C} > \text{C-H} > \text{H-H}$

Strength:  $\text{C-C} < \text{C-H} < \text{H-H}$

Partial positive charge:  $\text{C} > \text{H}$

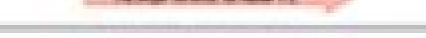
Partial negative charge:  $\text{H} > \text{C}$

#### Relative bond strength



Bond strength:  $\text{C-C} < \text{C-H} < \text{H-H}$

#### Relative bond to water



Bond to water:  $\text{C-C} < \text{C-H} < \text{H-H}$

# Organic Chemistry Study Guide For Pcat

**Robert R. Redfield**

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## **Table of Contents Organic Chemistry Study Guide For Pcat**

1. Understanding the eBook Organic Chemistry Study Guide For Pcat
  - The Rise of Digital Reading Organic Chemistry Study Guide For Pcat
  - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chemistry Study Guide For Pcat
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Organic Chemistry Study Guide For Pcat
  - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chemistry Study Guide For Pcat
  - Personalized Recommendations
  - Organic Chemistry Study Guide For Pcat User Reviews and Ratings
  - Organic Chemistry Study Guide For Pcat and Bestseller Lists
5. Accessing Organic Chemistry Study Guide For Pcat Free and Paid eBooks
  - Organic Chemistry Study Guide For Pcat Public Domain eBooks
  - Organic Chemistry Study Guide For Pcat eBook Subscription Services
  - Organic Chemistry Study Guide For Pcat Budget-Friendly Options
6. Navigating Organic Chemistry Study Guide For Pcat eBook Formats

- ePub, PDF, MOBI, and More
- Organic Chemistry Study Guide For Pcat Compatibility with Devices
- Organic Chemistry Study Guide For Pcat Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Organic Chemistry Study Guide For Pcat
- Highlighting and Note-Taking Organic Chemistry Study Guide For Pcat
- Interactive Elements Organic Chemistry Study Guide For Pcat

8. Staying Engaged with Organic Chemistry Study Guide For Pcat

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Organic Chemistry Study Guide For Pcat

9. Balancing eBooks and Physical Books Organic Chemistry Study Guide For Pcat

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Organic Chemistry Study Guide For Pcat

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Organic Chemistry Study Guide For Pcat

- Setting Reading Goals Organic Chemistry Study Guide For Pcat
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Organic Chemistry Study Guide For Pcat

- Fact-Checking eBook Content of Organic Chemistry Study Guide For Pcat
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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