

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Interactions

There are two types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



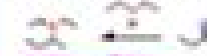
- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Strength of the



Bond and Bonding



Bonding energy



There are two types of

Definition of Polarity



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- A single bond is a sigma bond.
- A double bond consists of one sigma and one pi bond.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



The direction of the dipole moment is from the positive end to the negative end.

- The dipole moment is a vector pointing from the positive end to the negative end.
- The vector dipole moment is a vector pointing from the positive end to the negative end.
- μ (dipole moment) indicates an electron shift.
- δ (dipole moment) indicates an electron shift.

Molecular Dipole

To determine the molecular dipole, the vector sum of the dipole moments.

The polarity of the molecule is a vector sum of the dipole moments of the bonds.



Sum the vector of the dipole moments to find the net dipole.

Bond Length and Bond Strength

The length of the bond is a vector sum.



Double bonds are stronger than single bonds.

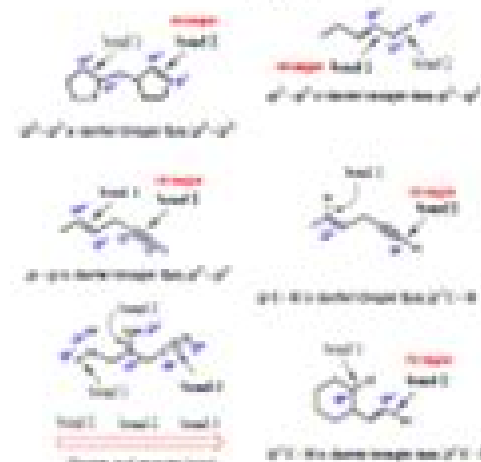


The length of the bond is a vector sum.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	112 kcal/mol	200 kcal/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.09 Å	1.08 Å	1.06 Å
Strength	98 kcal/mol	110 kcal/mol	120 kcal/mol

Higher s character = stronger and shorter bond

Examples



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Robert R. Redfield



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