

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium chlorides: $\text{NH}_4^+ + \text{H}_3\text{N}^- \rightarrow \text{NH}_3 + \text{NH}_4^+$



Four different types of bonds

1. Ionic bond
2. Covalent bond
3. Hydrogen bond
4. Metallic bond

There are four types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electron from one atom to another (anions and cations). (Opposite charges).



The covalent bonding of two atoms

$\text{O}=\text{C}=\text{O}$

Shared and transferred electrons

H_2

Hydrogen bonding

H_2O

Electron transfer

Na^+

Metallic bonding

Na^+

Hydrogen bonding

H_2O

Electron transfer

Na^+

Hydrogen bonding

H_2O

Electron transfer

Na^+

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Protophiles



Reactions and Reactions

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).

1. Nonpolar covalent bond
2. Polar covalent bond
3. Ionic bond



- The dipole moment is shown as a vector pointing from the less electronegative bond to the more electronegative atom.
- The vector is longer the larger the partial negative charge and the less electronegative atom is.
- The dipole bond is formed by a partial positive charge and a partial negative charge.
- The dipole bond is formed by a difference in electronegativity.
- The dipole bond is formed by a difference in electronegativity.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of the bond dipole moments. If the bond dipole moments are equal, it is just simple to calculate.



Sum of bond dipole moment = molecular dipole moment

Bond Length and Bond Strength

Reactions and Reactions



Bond length and bond strength

Strong bond = Short bond

Weak bond = Long bond

Strong bond = Short bond

Weak bond = Long bond

Strong bond = Short bond

Weak bond = Long bond

Strong bond = Short bond

Weak bond = Long bond

Strong bond = Short bond

Weak bond = Long bond

Strong bond = Short bond

Weak bond = Long bond

Strong bond = Short bond

Weak bond = Long bond

Reactions and Reactions

Average values of C-C and C-H bond depending on hybridization

C-C Bond	$\text{C}_2\text{H}_2\text{C}_2\text{H}_2$	$\text{C}_2\text{H}_4\text{C}_2\text{H}_4$	$\text{C}_2\text{H}_6\text{C}_2\text{H}_6$
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Hybridization	sp^2 (sp ²)	sp^3 (sp ³)	sp^3 (sp ³)
---------------	----------------------------------	----------------------------------	----------------------------------

Length	1.37 Å	1.35 Å	1.35 Å
--------	--------	--------	--------

Strength	1.01 kcal/mol	1.11 kcal/mol	1.01 kcal/mol
----------	---------------	---------------	---------------

Hybridization	sp^2	sp^3	sp^3
---------------	---------------	---------------	---------------

Hybridization	sp^2	sp^3	sp^3
---------------	---------------	---------------	---------------

Length	1.35 Å	1.35 Å	1.35 Å
--------	--------	--------	--------

Strength	1.01 kcal/mol	1.01 kcal/mol	1.01 kcal/mol
----------	---------------	---------------	---------------

Hybridization and bond length

Example



1. C_2H_2 (shortest bond) C_2H_2

2. C_2H_4 (medium bond) C_2H_4

3. C_2H_6 (longest bond) C_2H_6

1. C_2H_2 (shortest bond) C_2H_2

2. C_2H_4 (medium bond) C_2H_4

3. C_2H_6 (longest bond) C_2H_6

1. C_2H_2 (shortest bond) C_2H_2

2. C_2H_4 (medium bond) C_2H_4

3. C_2H_6 (longest bond) C_2H_6

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