

Organic chemistry - hydrocarbons

Week 1		
Crude oil Crude oil as a finite resource, remains of ancient biomass, mainly plankton. Mixture of a very large number of compounds, mostly hydrocarbons, made of hydrogen and carbon atoms only.	Alkanes General formula C_nH_{2n+2} Names of first 4 alkanes (methane, ethane, propane and butane). How alkanes are represented. Link to strong covalent bonds within molecules and weak intermolecular forces between molecules. Recognise alkanes. Know that alkanes are fully saturated.	Alkenes More reactive than alkanes. React with bromine water, which is used as a test for alkenes.
Week 2		
Properties of hydrocarbons Properties of hydrocarbons. Some properties of hydrocarbons depend on the size of their molecules, including boiling point, viscosity and flammability. These properties influence how hydrocarbons are used as fuels.	Fractional distillation MASTERY QUIZ 1 Fractional distillation and petrochemicals. How hydrocarbons may be separated into fractions with similar number of carbon atoms. Fractions used as fuels that we depend on and many other petrochemicals, such as solvents, lubricants, polymers and detergents.	Burning hydrocarbons Complete combustion of hydrocarbons. Combustion of fuel releases energy. The carbon and hydrogens are oxidised. Balance equations for the complete combustion of hydrocarbons with a given formula.
Week 3		
Cracking hydrocarbons Cracking. Hydrocarbons can be broken down to produce smaller, more useful molecules. High demand for fuel and how modern life depends on these. Catalytic cracking Steam cracking Products of cracking are alkanes and alkenes.	Polymers Polymers Alkenes are used to produce polymers and as starting materials for the production of many other chemicals.	Consolidation/exam practice OTT

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Arshad Iqbal

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