

- **Thermodynamics** is the study of energy and its transformations.
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.

$$\Delta U = Q - W$$

 where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.
 - **Second Law of Thermodynamics**: The entropy of an isolated system always increases over time.

$$\Delta S \geq 0$$

 where ΔS is the change in entropy.



- **Enthalpy** (H) is a thermodynamic property that represents the total energy of a system, including internal energy and the energy required to displace the environment.

$$H = U + PV$$

 where U is internal energy, P is pressure, and V is volume.
 - **Enthalpy Change** (ΔH) is the change in enthalpy during a process.

$$\Delta H = H_{\text{final}} - H_{\text{initial}}$$

 - **Exothermic Process**: A process that releases heat to the surroundings, resulting in a negative ΔH .
 - **Endothermic Process**: A process that absorbs heat from the surroundings, resulting in a positive ΔH .



- **Free Energy** (G) is a thermodynamic property that represents the maximum useful work that can be extracted from a system at constant temperature and pressure.

$$G = H - TS$$

 where H is enthalpy, T is temperature, and S is entropy.
 - **Free Energy Change** (ΔG) is the change in free energy during a process.

$$\Delta G = \Delta H - T\Delta S$$

 - **Spontaneous Process**: A process that occurs without the need for external energy input, resulting in a negative ΔG .
 - **Non-spontaneous Process**: A process that requires external energy input to occur, resulting in a positive ΔG .



- **Heat Capacity** (C) is the amount of heat required to raise the temperature of a substance by one degree Celsius.

$$Q = C\Delta T$$

 where Q is the heat added, C is the heat capacity, and ΔT is the change in temperature.
 - **Specific Heat Capacity** (c) is the heat capacity per unit mass of a substance.

$$Q = mc\Delta T$$

 where m is the mass of the substance.



- **Chemical Equilibrium** is a state in which the concentrations of reactants and products remain constant over time.

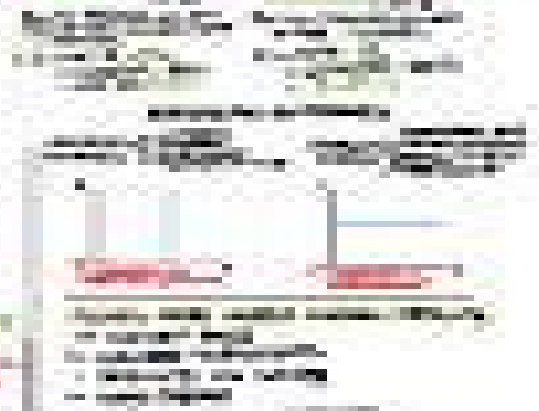
$$K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

 where K is the equilibrium constant, $[A]$, $[B]$, $[C]$, and $[D]$ are the concentrations of the reactants and products, and a , b , c , and d are the stoichiometric coefficients.
 - **Le Chatelier's Principle**: If a system at equilibrium is subjected to a change in concentration, pressure, or temperature, the system will shift its equilibrium position to counteract the change.
 - **Reaction Quotient** (Q) is a measure of the relative amounts of reactants and products in a reaction mixture.

$$Q = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

 - **Equilibrium Constant** (K) is a measure of the extent to which a reaction proceeds towards equilibrium.

$$K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$



Microeconomics Final Study Guide

Edward Tower



Microeconomics Final Study Guide:

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CLEP Official Study Guide College Entrance Examination Board, 2006-08 A guide to preparing for the thirty four different kinds of college CLEP exams that includes information on deciding which exams to take interpreting the scores taking the tests and other related topics

Resources in Education, 1980

Official Study Guide for the CLEP Examinations, 1998 College Entrance Examination Board, 1997 The College Level Examination Program CLEP tests are comprehensive 90 minute exams used to award full college credit for demonstrating college level achievement in a variety of areas in subjects This guide helps students decide which CLEP exams to take helps them interpret their CLEP scores and includes sample questions and answers for all 34 exams and a list of study resources

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This official guide developed by the sponsors of the CLEP exam includes a CD ROM for Windows that introduces test takers to the new computer based test and provides a tutorial for each exam Study Guide for Use with Microeconomics, Second Canadian Edition David C. Colander, 2003 **Study Guide for Baumol and Blinder's Economics, Principles and Policy** Craig Swan, 1985 *Study Guide for Baumol and Blinder's Economics, Principles and Policy, Fifth Edition* Craig Swan, 1991

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Economics Gary W. Yohe, 1992

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