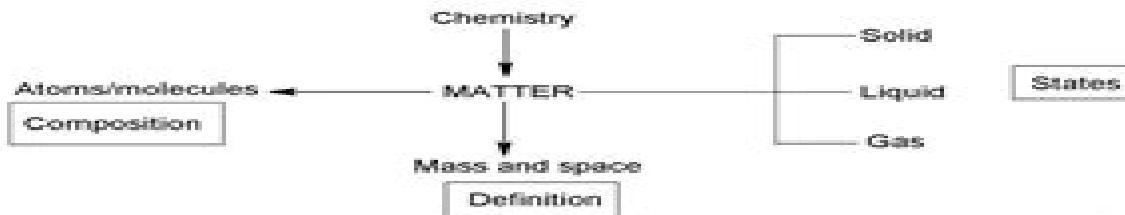


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**Atoms:** Building block of matter. It may or may not exist free in nature

**Examples:** H, He, Na, C

Shells = electrons



Nucleus = Protons + Neutrons

Mass = 0

Charge = +1 0

Charge = -1

Mass = 1 1

**Nucleus:** It is central part of an atom.

**Shells:** Circular paths around the nucleus where electrons reside. They are also called orbits or energy levels.

No. of es in a shell =  $2n^2$  Where n = no of shell.

1<sup>st</sup> shell = 2es, 2<sup>nd</sup> shell = 8es, 3<sup>rd</sup> shell = 18es.

At least 8 electron should be present in last for the stability of an atom. However if an atom has one shell like Hydrogen and Helium then 2 electrons are required for complete shell and stability.

**Atomic number:**

Number of protons present in the nucleus of an atom of an element.

Examples: 11Na, 6C, 8O

**Atomic mass or Mass number or nucleon number:** It is the sum of protons and neutrons present in the nucleus of an atom of an element.

Examples: <sup>23</sup>Na, <sup>12</sup>C, <sup>16</sup>O

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**Léo Pomar**

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