

# NCERT Solutions for Class-XI Chemistry

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## Chapter-2

### NCERT Chemistry Class 11

1. Calculate the molecular mass of the following:

- Calculate the number of electrons which will together weigh one gram.
- Calculate the mass and charge of one mole of electrons.

(i) Mass of one electron =  $9.10939 \times 10^{-31}$  kg

∴ Number of electrons that weigh  $9.10939 \times 10^{-31}$  kg = 1

∴ Number of electrons that will weigh 1 g ( $1 \times 10^{-3}$  kg)

$$= \frac{1}{9.10939 \times 10^{-31} \text{ kg}} \times (1 \times 10^{-3} \text{ kg})$$
$$= 0.1098 \times 10^{31} + 3.1$$
$$= 0.1098 \times 10^{28}$$
$$= 1.098 \times 10^{27}$$

(ii) Mass of one electron =  $9.10939 \times 10^{-31}$  kg

Mass of one mole of electron =  $(6.022 \times 10^{23}) \times (9.10939 \times 10^{-31} \text{ kg})$   
=  $5.48 \times 10^{-7}$  kg

Charge on one electron =  $1.6022 \times 10^{-19}$  coulomb

Charge on one mole of electron =  $(1.6022 \times 10^{-19} \text{ C}) (6.022 \times 10^{23})$   
=  $9.65 \times 10^4$  C

2. (i) Calculate the total number of electrons present in one mole of methane.

(ii) Find (a) the total number and (b) the total mass of neutrons in 7 mg of  $^{14}\text{C}$ . (Assume that mass of a neutron =  $1.675 \times 10^{-27}$  kg).

(iii) Find (a) the total mass of protons in 34 mg of  $\text{NH}_3$  at STP.

Will the answer change if the temperature and pressure are changed?

2. No the answer will not vary with the change in temperature and pressure because the number of subatomic particles like protons, neutrons and electrons is fixed for each and every element and it does not vary with temperature and pressure.

3. How many neutrons and protons are there in the following nuclei?

$^{12}\text{C}$ ,  $^{16}\text{O}$ ,  $^{24}\text{Mg}$ ,  $^{56}\text{Fe}$ ,  $^{88}\text{Sr}$

3.  $^{13}\text{C}$ : Atomic mass = 13

Atomic number = Number of protons = 6

Number of neutrons = (Atomic mass) - (Atomic number)

$$= 13 - 6 = 7$$

$^{16}\text{O}$ :

Atomic mass = 16

Atomic number = 8

Number of protons = 8

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**Lakhmir Singh & Manjit Kaur**

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### **Table of Contents Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com**

1. Understanding the eBook Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
  - The Rise of Digital Reading Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
  - Advantages of eBooks Over Traditional Books
2. Identifying Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
  - Personalized Recommendations
  - Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com User Reviews and Ratings
  - Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com and Bestseller Lists

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  - ePUB, PDF, MOBI, and More
  - Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com Compatibility with Devices
  - Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
  - Highlighting and Note-Taking Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
  - Interactive Elements Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
8. Staying Engaged with Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
9. Balancing eBooks and Physical Books Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
  - Setting Reading Goals Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
  - Fact-Checking eBook Content of Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
  - Distinguishing Credible Sources
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- Exploring Educational eBooks

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- Interactive and Gamified eBooks

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