



Overview Chemical Bonds

Directions: All of the statements below are *false* as written. In the space provided, write a term or phrase that makes the statement *true* when it is substituted for the underlined words.

1. The properties of a compound are the same as the properties of the elements that it contains.
2. Superscript numbers in chemical formulas tell how many atoms of each element are found in a unit of compound.
3. All the noble gases except helium have 18 electrons in their outer energy level.
4. A(n) chemical formula is the force that holds atoms together in a compound.
5. An ion is a(n) neutral particle that has either more or fewer electrons than protons.
6. Oxidation numbers are written as subscripts.
7. A(n) covalent bond is the force of attraction between the opposite charges of the ions in an ionic compound.
8. The charge on a compound is always positive.
9. Equal sharing of electrons in covalent bonds results in polar molecules.
10. Only two identical atoms can share electrons unequally.
11. A binary compound contains five different elements.
12. An oxidation number tells how many protons an atom must gain, lose, or share to become stable.
13. The oxidation number of the copper(II) ion is 3+.
14. When writing chemical formulas, add superscripts so that the sum of the oxidation numbers equals ten.
15. A polyatomic ion never has a positive or negative charge.
16. The polyatomic ion SO_4^{2-} is called the sulfide ion.

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