

CHAPTER **11** STUDY GUIDE

Stoichiometry

Section 11.1 What is stoichiometry?

In your textbook, read about stoichiometry and the balanced equation.

For each statement below, write *true* or *false*.

- _____ 1. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.
- _____ 2. Stoichiometry is based on the law of conservation of mass.
- _____ 3. In any chemical reaction, the mass of the products is less than the mass of the reactants.
- _____ 4. The coefficients in a chemical equation represent not only the number of individual particles but also the number of moles of particles.
- _____ 5. The mass of each reactant and product is related to its coefficient in the balanced chemical equation for the reaction by its molar mass.

Complete the table below, using information represented in the chemical equation for the combustion of methanol, an alcohol.



Substance	Molar Mass (g/mol)	Number of Molecules	Number of Moles (mol)	Mass (g)
6. Methanol	32.05			
7. Oxygen gas	32.00			
8. Carbon dioxide	44.01			
9. Water	18.02			

10. What are the reactants? _____
11. What are the products? _____
12. What is the total mass of the reactants? _____
13. What is the total mass of the products? _____
14. How do the total masses of the reactants and products compare? _____

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1. Understanding the eBook Mcgraw Hill Study Guide Frankenstein
 - The Rise of Digital Reading Mcgraw Hill Study Guide Frankenstein
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Study Guide Frankenstein
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Study Guide Frankenstein
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Study Guide Frankenstein
 - Personalized Recommendations
 - Mcgraw Hill Study Guide Frankenstein User Reviews and Ratings
 - Mcgraw Hill Study Guide Frankenstein and Bestseller Lists
5. Accessing Mcgraw Hill Study Guide Frankenstein Free and Paid eBooks
 - Mcgraw Hill Study Guide Frankenstein Public Domain eBooks
 - Mcgraw Hill Study Guide Frankenstein eBook Subscription Services
 - Mcgraw Hill Study Guide Frankenstein Budget-Friendly Options
6. Navigating Mcgraw Hill Study Guide Frankenstein eBook Formats

- ePub, PDF, MOBI, and More
- Mcgraw Hill Study Guide Frankenstein Compatibility with Devices
- Mcgraw Hill Study Guide Frankenstein Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcgraw Hill Study Guide Frankenstein
 - Highlighting and Note-Taking Mcgraw Hill Study Guide Frankenstein
 - Interactive Elements Mcgraw Hill Study Guide Frankenstein
- 8. Staying Engaged with Mcgraw Hill Study Guide Frankenstein
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Study Guide Frankenstein
- 9. Balancing eBooks and Physical Books Mcgraw Hill Study Guide Frankenstein
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Study Guide Frankenstein
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mcgraw Hill Study Guide Frankenstein
 - Setting Reading Goals Mcgraw Hill Study Guide Frankenstein
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mcgraw Hill Study Guide Frankenstein
 - Fact-Checking eBook Content of Mcgraw Hill Study Guide Frankenstein
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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