

REDOX REACTIONS.

NOTES...

Intro

Many Chemical reactions involve transfer of electrons from one chemical substance to another. These electron-transfer reactions are termed as Oxidation-Reduction or Redox reactions.

Oxidation

(i) Loss of electron

- Example: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$

- Sub-points:

- Loss of hydrogen
- Gain of oxygen
- Gain of **electronegative** element OR loss of electropositive element

Reduction

i) Gain of electron

- Example: $\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$

- Sub-points:

- Gain of hydrogen
- Loss of oxygen
- Loss of electronegative element OR gain of electropositive element

Redox reaction

The reaction in which oxidation as well as reduction takes place together is known as a redox reaction.

REDOX REACTIONS IN TERMS OF ELECTRON TRANSFER (MODERN IDEA)

Oxidation process is the process in which one or more electrons are lost by an atom, ion, or molecule.

Example: $\text{Zn} \rightarrow \text{Zn}^{2+} + 2\text{e}^-$

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John McMurry, Susan McMurry

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Table of Contents Mcgraw Hill Oxidation Reduction Study Guide

1. Understanding the eBook Mcgraw Hill Oxidation Reduction Study Guide
 - The Rise of Digital Reading Mcgraw Hill Oxidation Reduction Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Oxidation Reduction Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Oxidation Reduction Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Oxidation Reduction Study Guide
 - Personalized Recommendations
 - Mcgraw Hill Oxidation Reduction Study Guide User Reviews and Ratings

- Mcgraw Hill Oxidation Reduction Study Guide and Bestseller Lists
- 5. Accessing Mcgraw Hill Oxidation Reduction Study Guide Free and Paid eBooks
 - Mcgraw Hill Oxidation Reduction Study Guide Public Domain eBooks
 - Mcgraw Hill Oxidation Reduction Study Guide eBook Subscription Services
 - Mcgraw Hill Oxidation Reduction Study Guide Budget-Friendly Options
- 6. Navigating Mcgraw Hill Oxidation Reduction Study Guide eBook Formats
 - ePUB, PDF, MOBI, and More
 - Mcgraw Hill Oxidation Reduction Study Guide Compatibility with Devices
 - Mcgraw Hill Oxidation Reduction Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcgraw Hill Oxidation Reduction Study Guide
 - Highlighting and Note-Taking Mcgraw Hill Oxidation Reduction Study Guide
 - Interactive Elements Mcgraw Hill Oxidation Reduction Study Guide
- 8. Staying Engaged with Mcgraw Hill Oxidation Reduction Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Oxidation Reduction Study Guide
- 9. Balancing eBooks and Physical Books Mcgraw Hill Oxidation Reduction Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Oxidation Reduction Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mcgraw Hill Oxidation Reduction Study Guide
 - Setting Reading Goals Mcgraw Hill Oxidation Reduction Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mcgraw Hill Oxidation Reduction Study Guide
 - Fact-Checking eBook Content of Mcgraw Hill Oxidation Reduction Study Guide
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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