

9.1. INTRODUCTION

- Potentiometric titration embrace two major types of analyses
- It involves the direct measurement of an electrode potential from which the activity (or concentration) of an active ion may be derived.
- The second type involves measuring the changes in electromotive force (EMF) brought about by the addition of a titrant to the sample.

9.1.1. PRINCIPLE

- The principle involved in the potentiometry is when the pair of electrodes is placed in the sample solution it shows the potential difference by the addition of the titrant or by the change in the concentration of the ions. This can be seen below:

1. Reference electrode/salt bridge/analyte solution/indicator electrode Potentiometer

- The reference electrode is the electrode which consists of its own potential value and it is stable when dipped into the sample solution.
- The salt bridge is used to prevent the interference of the analyte solution with the reference solution.
- Analyte solution is the sample solution whose potential is to be measured.

2. The indicator electrode is the electrode which responds to the changes in the potential of the analyte solution.

The electromotive force of the complete cell is given by the following equation:

$$E_{\text{cell}} = E_{\text{reference}} + E_{\text{indicator}} + E_{\text{junction}}$$

Where $E_{\text{reference}}$ is the electromotive force of the reference electrode;

$E_{\text{indicator}}$ is the electromotive force of the indicator electrode; E_{junction} is the electromotive force at the junction of the liquid.

- The units for the potential are mille volts denoted by mV.

THEORY

- The main theory involved in the potentiometry is when the known potential electrode immersed in the sample solution then the potential is given by the Nernst equation:

$$E = E^{\circ} + (0.592/n) \log c$$

where

- E is the potential of the solution;
- E° is the standard electrode potential;
- n is the valency of the ions;
- c is the concentration of the sample solution;

Mcq On Potentiometry

G. Svehla

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Advanced Potentiometry Erzsébet Néher-Neumann,2010-03-14 Advanced

Potentiometry details the accurate calculation of potentiometric titrations Additional terms such as the ideal diffusion potential together with the contribution of activity factors furnish calculated stability constants which do not include systematic errors and thus present a complete and correct description of equilibrium processes Moreover serious systematic errors resulting from the recent use of curve fitting computer programs are highlighted later in the text Advanced Potentiometry can advantageously be used by lecturers researchers PhD graduates and undergraduate students working in the fields of Solution Chemistry Physical Chemistry Inorganic Chemistry and Analytical Chemistry and Environmental Research

Improved Equipment for Potentiometric Titration Procedure Edward N. Wise,1950 **Class 11-12**

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