

# Study Guide Unit 1-2: Matter and Measurements



1. Classify each of the following as an element, compound, homogeneous or heterogeneous mixture.

(a) air \_\_\_\_\_ (b) tomato juice \_\_\_\_\_  
(c) iodine (I<sub>2</sub>) crystals \_\_\_\_\_ (d) sand (NaCl) \_\_\_\_\_

2. (a) Is a mixture of sugar and sand homogeneous or heterogeneous? \_\_\_\_\_

(b) Explain how you determined the answer to part (a). \_\_\_\_\_

3. In 1807, the English chemist Humphry Davy passed an electric current through molten potassium hydroxide and isolated a bright, shiny, conductive substance.

a. Would you classify the substance as a metal, nonmetal or metalloid? \_\_\_\_\_  
b. Was the change that occurred chemical or physical? Explain. \_\_\_\_\_

4. Read the following description of the element magnesium. Circle the physical properties and underline the chemical properties.

Magnesium is a silver-gray colored metal with a melt point of 650°C. When magnesium is placed in acid, it reacts effervescently releasing hydrogen gas. Magnesium has a flammability on the Rober's scale of 2.5 and a density of 1.74 g/cm<sup>3</sup> at 25°C. It reacts with oxygen gas when heated to form zinc oxide, MgO.

5. A match is lit and held under a cold piece of metal. The following observations are made:

(a) The match burns. (b) The metal gets warmer. (c) Water condenses on the metal. (d) Soot (carbon) is deposited on the metal.

Which letters exhibit occurrences of physical changes? \_\_\_\_\_ Chemical changes? \_\_\_\_\_

6. Determine if A-E are observations or interpretations: A match is lit and held under a cold piece of metal. The following occurs.

a. The match burns. \_\_\_\_\_ b. The match reacts with oxygen. \_\_\_\_\_  
c. The metal gets warmer. \_\_\_\_\_ d. Heat transfers to the metal from the match. \_\_\_\_\_  
e. Black soot is deposited on the metal. \_\_\_\_\_

7. Circle the larger unit: a. centigram or milligram b. liter or kiloliter c. millisecond or microsecond

8. Fill in the blank (Know the following):

a. 1 cm = \_\_\_\_\_ mm  
b. 1 km = \_\_\_\_\_ m  
c. 1 m = \_\_\_\_\_ cm  
d. 1 m = \_\_\_\_\_ cm

e. 1 L = \_\_\_\_\_ ml  
f. 1 kg = \_\_\_\_\_ g  
g. 1 g = \_\_\_\_\_ mg  
h. 1 g = \_\_\_\_\_ kg

i. 1 mm = \_\_\_\_\_ m  
j. 1 m = \_\_\_\_\_ mm  
k. 1 mm = \_\_\_\_\_ m  
l. 1 m = \_\_\_\_\_ mm

# Measurements And Calculations Chemistry Study Guide

## Answers

**Darrell D. Ebbing**

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**Study Guide to Chemistry and Life** John William Hill, Dorothy M. Feigl, 1978

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