

## **Metals in Aqueous Solutions Simulations**

### **How to use it:**

Some elements are more reactive than others, meaning that they can more easily be broken apart into their ions than others. These are listed in their place respectively in "More" boxes. Other elements are less reactive and more likely not to be broken as easily. And in a comparison. This simulation will allow you to test various combinations of elements and compounds, as well as reactions of elements, to see how they react together and how much of each is left or left over. **Activity Series**. Some metals are more willing to lose electrons than others, as a compound gives it up. For example, for example, some metals, including all metals, are more likely to be broken apart with compounds, including more reactive elements.

### **Objectives**

Use the results of the simulation program to determine the activity series for the elements using this simulation.

### **Procedure**

For this, the simulation at: <https://chemistry.mhhe.com/chem101/chem101sim/01/01sim01.html>, <https://chemistry.mhhe.com/chem101/chem101sim/01/01sim01.html> and complete parts of the **PCBP**, including **Activity 1.1**, and the **Observation** section, recording the appropriate information and answering the questions below.

The screenshot shows the 'Metals in Aqueous Solutions' simulation interface. On the left, there is a title bar and a description: 'Investigate the reactivity of metals in aqueous solutions.' Below this is a 'Start' button. On the right, there is a grid of six beakers, each containing a different colored liquid (blue, green, yellow, orange, red, and purple) and a metal strip. The beakers are labeled with metal names: Aluminum, Zinc, Iron, Lead, Silver, and Gold. The simulation is set to 'Standard Conditions' and 'Aqueous Solution'.

# Metals In Aqueous Solutions Answers

**Priti Singhal**



## Metals In Aqueous Solutions Answers:

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