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mole H_2S x 18g H_2O i mole H_2O 149 4g H_2O is your theoretical yield so percent yield $\frac{137}{149} \times 100 = 91.77\%$ problems 6 8 are similar to problem 5

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al x $\frac{1 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} \times 101.964 \text{ g Al}_2\text{O}_3 = 81.58 \text{ g Al}_2\text{O}_3$ c determine the percent yield $\frac{66.0}{81.58} \times 100 = 80.9\%$

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