

# Reactions of Amines and Epoxides

## Preparation of Amines

Amines are often made by reduction of nitriles or by Gabriel phthalimide synthesis



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



## Bond Types, Strength, Length, Polarity and Dipole Moment

### Types of Bonding

There are two types of bonding: ionic bonding and covalent bonding

- Ionic bonds are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron gain enthalpy).



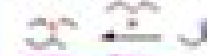
- Covalent bonds are formed by sharing of electrons between two atoms.



### The Relative Strength of Ions



### Bond and Bonding



### Bonding



### There are two types of bonding

Definition of ionic and covalent bonds



## Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds

- Sigma bonds are formed by head-on (end-to-end) overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- A single bond is a sigma bond.
- A double bond consists of one sigma and one pi bond.



pi bonds are weaker than sigma bonds.

## Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond) or a lower bond is formed.

## Dipole Moment

The polarity of a bond is described by the dipole moment ( $\mu$ ).



- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge ( $\delta^-$ ) and the less electronegative atom has a partial positive charge ( $\delta^+$ ).
- $\delta^-$  (delta minus) indicates an electron deficient atom.
- $\delta^+$  (delta plus) indicates an electron rich atom.

## Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipole moments are present. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipoles cancel out.

## Bond Length and Bond Strength

The length of a bond is important.

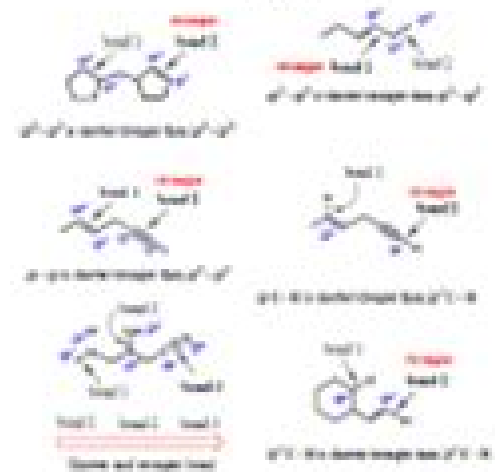


The length of a bond is important.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	$sp^3-sp^3$ (C-C)	$sp^2-sp^2$ (C=C)	$sp-sp$ (C $\equiv$ C)
Hybridization	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	65 kcal/mol	200 kcal/mol
Hybridization	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
C-H Bond	1.09 Å	1.08 Å	1.06 Å
Strength	98 kcal/mol	110 kcal/mol	120 kcal/mol

Higher s character means stronger bond

## Examples



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**Robert J. Ouellette, J. David Rawn**



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