

Reactions of Amines and Epoxides

Preparation of Amines

Amines are often made by reduction of nitriles or by reductive amination of aldehydes and ketones.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity. Examples:



• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bonded and Unbonded electron pairs



Bonding electrons



There are two main categories

1. Definition of electron pair



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

• **Sigma bonds** are formed by head-on overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **One single bond is a sigma bond.**

• **Two double bonds consist of one sigma and one pi bond.**



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.

• **Polar bonds** are formed by the difference in electronegativity values of the two atoms, forming a dipole moment.

• **Nonpolar bonds** are formed by the same electronegativity values of the two atoms, forming no dipole moment.

• If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared equally (nonpolar) or the electrons are shared unequally (polar).

• If the difference in electronegativity is between 1.7 and 2.0, the electrons are shared unequally (polar) or the electrons are shared equally (nonpolar).

• A difference of 2.0 or higher indicates that the electrons are not shared (ionic bond) or the bond is ionic.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



Electronegative atoms are color-coded to represent the dipole moment.

• The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.

• The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).

• **Electronegativity** indicates an electron's ability to attract electrons.

• **Electronegativity** indicates an electron's ability to attract electrons.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is used.

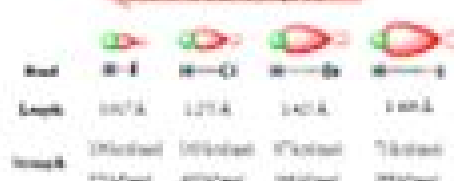
The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipole moments are non-zero. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipole moment is zero.

Bond Length and Bond Strength

The length of the bond is important.



Double bonds are stronger than single bonds and are shorter.

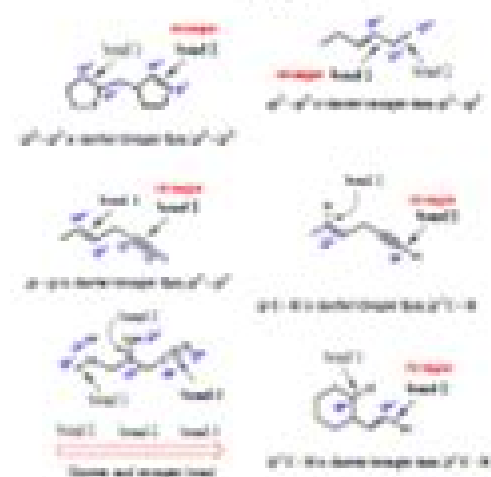


The length of the bond is important.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.09 Å	1.08 Å	1.06 Å
Strength	98 kcal/mol	110 kcal/mol	120 kcal/mol

Higher s character means stronger bond

Examples



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