

# Reactions of Amines and Epoxides

## Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for tertiary hydroxyl group



A Summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



## Bond Types, Strength, Length, Polarity and Dipole Moment

### Two types of covalent bond

There are two types of covalent bond: bonding and non-bonding bond

- **Bonding bonds** are formed by the overlap of electron orbitals and are shared between two atoms (positive) and **antibonding** (negative) bonds



- **Covalent bonds** are formed by sharing of electrons between two atoms



### The Relative Strength of $\sigma$ and $\pi$ Bonds



### Bonding and Antibonding Orbitals



### Bonding energy



### There are two types of covalent bond

Definition of covalent bond



## Sigma ( $\sigma$ ) and pi ( $\pi$ ) bonds

Covalent bonds are classified into sigma ( $\sigma$ ) and pi ( $\pi$ ) bonds

- **Sigma bonds** are formed by head-on overlap of orbitals
- **pi bonds** are formed by side-on overlap of orbitals
- **One single bond is a sigma bond**
- **One double bond consists of one sigma and one pi bond**

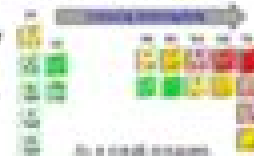


pi bonds are stronger than sigma bonds

## Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity

covalent bonds can be polar and nonpolar



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond), forming a **nonpolar covalent bond**. Two important examples are the C-C and C-H bonds.

- If the difference in electronegativity is between 0.5 and 1.7, we have a **polar covalent bond**. Three important examples are bonds for carbon to fluorine, N, O, P and other halogens.

- A difference of 1.7 or higher can be high that the electrons are not shared (ionic bond) and a **ionic bond** is formed.

## Dipole Moment

The polarity of a bond is described by the **dipole moment** ( $\mu$ )



- The **dipole moment** is defined as a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge ( $\delta^-$ ) and the less electronegative atom has a partial positive charge ( $\delta^+$ ).
- **$\delta^-$  (delta minus)** indicates an electron deficient atom.
- **$\delta^+$  (delta plus)** indicates an electron rich atom.

## Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used

The geometry of the molecule is important. Symmetrical molecules are nonpolar (even for dipole moments are cancelled). A good example is carbon dioxide.



Cancel the vector or the other pointing point in the same way

## Bond Length and Bond Strength

The longer the bond, the weaker the bond

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds in the same row and a stronger bond

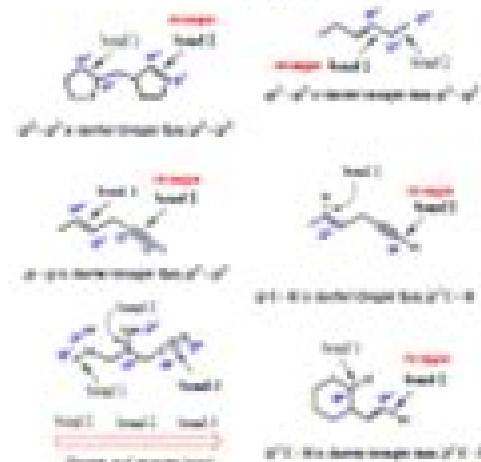


The longer the bond, the weaker the bond

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	$sp^3-sp^3$ C-C	$sp^2-sp^2$ C-C	$sp-sp$ C-C
Hybridization	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	611 kJ/mol	837 kJ/mol
Hybridization	$sp^3-sp^2$	$sp^2-sp$	$sp-sp$
C-H Bond	109 pm	108 pm	106 pm
Strength	413 kJ/mol	439 kJ/mol	505 kJ/mol

Higher s character = stronger and shorter bond

## Examples



gaining information from the graph

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# Organic Chemistry Final Study Guide

**Marye Anne Fox, James K. Whitesell**



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