

# Reactions of Amines and Epoxides

## Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

## Bond Types, Strength, Length, Polarity and Dipole Moment

### Free-radical halogenation



There are two types of halogen bond: covalent and non-covalent.

- **Covalent bonds** are formed by the sharing of electron pairs between two atoms (positive and negative charges).



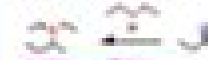
- **Non-covalent bonds** are formed by the sharing of electron pairs between two atoms.



### The Relative Strength of C-H



### Bond and Reaction Rates



### Bond Energy



### There are two types of halogen bond

Definition of halogen bond



## Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

- **Sigma bonds** are formed by the end-to-end overlap of orbitals.

• **Pi bonds** are formed by the side-to-side overlap of orbitals.

- **Single bonds** are sigma bonds.

- **Double bonds** are composed of one sigma and one pi bond.



• **Double bonds** are stronger than single bonds.

## Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity:

• **Polar covalent bonds** are formed between atoms with different electronegativities.

• **Nonpolar covalent bonds** are formed between atoms with similar electronegativities.

• **Electronegativity** is a measure of an atom's ability to attract electrons.

• **Electronegativity difference** is the difference in electronegativity between two atoms.

• **Electronegativity difference** is a measure of the polarity of a bond.

• **Electronegativity difference** is a measure of the strength of a bond.

• **Electronegativity difference** is a measure of the stability of a bond.

• **Electronegativity difference** is a measure of the reactivity of a bond.

• **Electronegativity difference** is a measure of the toxicity of a bond.

• **Electronegativity difference** is a measure of the flammability of a bond.

• **Electronegativity difference** is a measure of the explosiveness of a bond.

• **Electronegativity difference** is a measure of the radioactivity of a bond.

## Dipole Moment

The polarity of a bond is described by the dipole moment ( $\mu$ ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

- The vector dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

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## Molecular Dipole

To determine the molecular dipole, the vector sum of the bond dipoles is used.

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Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
Hybridization	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	376 kJ/mol	611 kJ/mol	346 kJ/mol
Hybridization	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
C-H Bond	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
Length	1.09 Å	1.08 Å	1.06 Å
Strength	413 kJ/mol	505 kJ/mol	548 kJ/mol

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**Daniel Evan Jones**



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