

Specific Heat Problems – ANSWERS

The last page of this document has Miscellaneous Information that you will need for these problems.

1. How many calories are needed to warm 55.8 g of water at 21.3°C to 73.4°C?

$$\text{Heat} = m \times S. \text{H.} \times (T_{\text{final}} - T_{\text{initial}})$$

$$\text{Heat} = 55.8 \text{ g} \times 1.00 \text{ cal/g°C} \times (73.4^\circ\text{C} - 21.3^\circ\text{C})$$

$$\text{Heat} = 55.8 \text{ g} \times 1.00 \text{ cal/g°C} \times 52.1^\circ\text{C}$$

$$\begin{aligned}\text{Heat} &= 2907.18 \text{ cal} \text{ convert to scientific notation} \\ &= 2.90718 \times 10^3 \text{ cal} \\ &= 2.91 \times 10^3 \text{ cal (3 sig. figs.)}\end{aligned}$$

2. How many calories are needed to warm 55.8 g of alcohol from 21.3°C to 73.4°C?

$$\text{Heat} = m \times S. \text{H.} \times (T_{\text{final}} - T_{\text{initial}})$$

$$\text{Heat} = 55.8 \text{ g} \times 0.58 \text{ cal/g°C} \times (73.4^\circ\text{C} - 21.3^\circ\text{C})$$

$$\text{Heat} = 55.8 \text{ g} \times 0.58 \text{ cal/g°C} \times 52.1^\circ\text{C}$$

$$\begin{aligned}\text{Heat} &= 1686.16 \text{ cal} \text{ convert to scientific notation} \\ \text{Heat} &= 1.68616 \times 10^3 \text{ cal} \\ \text{Heat} &= 1.69 \times 10^3 \text{ cal}\end{aligned}$$

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Menu (Firm) (Fort Collins, Colo.)

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Unveiling the Energy of Verbal Artistry: An Mental Sojourn through **Organic Chemistry Specific Heat Problems**

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