

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Bonding Interactions

There are two types of bonding: ionic bonding and covalent bonding.

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of Ions



Bond and Bonding Interactions



Bonding Interactions



There are two types of bonding interactions



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

- **Sigma bonds** are formed by head-on overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **Single bonds** are sigma bonds.
- **Double bonds** are sigma and pi bonds.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic).

- A difference of 1.7 or higher can be used to determine if a bond is ionic or covalent.

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Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).
- δ⁺ (delta plus) indicates a partial positive charge.
- δ⁻ (delta minus) indicates a partial negative charge.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipoles cancel.

Bond Length and Bond Strength

The longer the bond, the weaker it is.

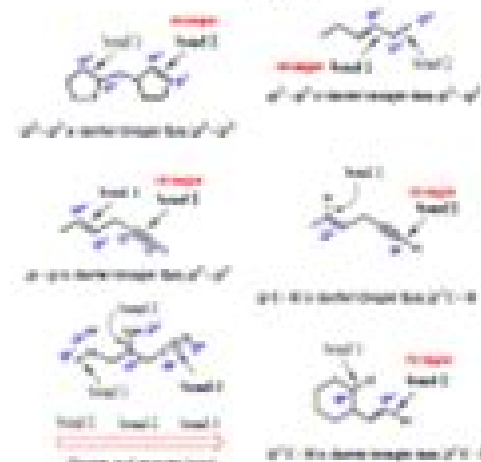


The longer the bond, the weaker it is.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp ³ -C-C	sp ² -C-C	sp-C-C
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp-sp
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp-sp
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol

Higher s character = stronger and shorter bond

Examples



Organic Chemistry Reactions Study Guide

T. W. Graham Solomons



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Table of Contents Organic Chemistry Reactions Study Guide

1. Understanding the eBook Organic Chemistry Reactions Study Guide
 - The Rise of Digital Reading Organic Chemistry Reactions Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chemistry Reactions Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chemistry Reactions Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chemistry Reactions Study Guide
 - Personalized Recommendations
 - Organic Chemistry Reactions Study Guide User Reviews and Ratings
 - Organic Chemistry Reactions Study Guide and Bestseller Lists

5. Accessing Organic Chemistry Reactions Study Guide Free and Paid eBooks
 - Organic Chemistry Reactions Study Guide Public Domain eBooks
 - Organic Chemistry Reactions Study Guide eBook Subscription Services
 - Organic Chemistry Reactions Study Guide Budget-Friendly Options
6. Navigating Organic Chemistry Reactions Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Organic Chemistry Reactions Study Guide Compatibility with Devices
 - Organic Chemistry Reactions Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Organic Chemistry Reactions Study Guide
 - Highlighting and Note-Taking Organic Chemistry Reactions Study Guide
 - Interactive Elements Organic Chemistry Reactions Study Guide
8. Staying Engaged with Organic Chemistry Reactions Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Organic Chemistry Reactions Study Guide
9. Balancing eBooks and Physical Books Organic Chemistry Reactions Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Organic Chemistry Reactions Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Organic Chemistry Reactions Study Guide
 - Setting Reading Goals Organic Chemistry Reactions Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Organic Chemistry Reactions Study Guide
 - Fact-Checking eBook Content of Organic Chemistry Reactions Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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