

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for indirect synthesis of amines



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: bonding and non-bonding bond

- Bonding bonds are formed by the overlap of electron orbitals and are classified as sigma bonds (single bonds) and pi bonds (double bonds).



- Non-bonding bonds are formed by the overlap of electron orbitals between two atoms.



The Relative Strength of Covalent Bonds



Bond and Bond Energy



Bond Energy



There are two types of covalent bond

Definition of covalent bond



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Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is defined as a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector electronegativity of an atom is a measure of its ability to attract electrons, and the dipole moment is a measure of the overall polarity of a molecule.
- δ^- (delta minus) indicates an electron-deficient atom.
- δ^+ (delta plus) indicates an electron-rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is calculated.

The polarity of the molecule is determined by the vector sum of the dipole moments. A polar molecule has a net dipole moment, while a non-polar molecule has no net dipole moment.



Calculate the vector sum of the dipole moments to determine the molecular dipole moment.

Bond Length and Bond Strength

The length of a bond is determined by the overlap of the atomic orbitals.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds, and triple bonds are stronger than double bonds.

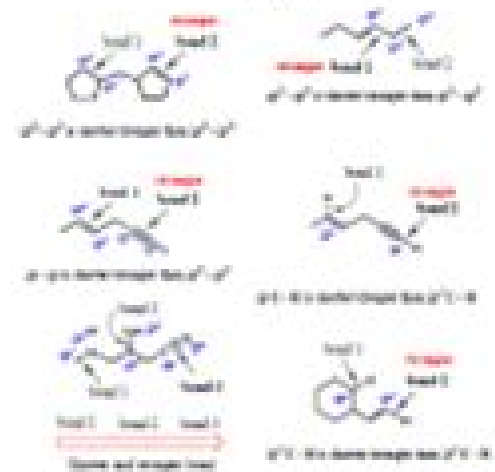


The length of a bond is determined by the overlap of the atomic orbitals.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	472 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
C-H Bond	109 pm	108 pm	106 pm
Length	109 pm	108 pm	106 pm
Strength	413 kJ/mol	479 kJ/mol	505 kJ/mol

Higher s character = stronger and shorter bond

Examples



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T. W. Graham Solomons



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Study Guide with Solutions to Accompany Burgoyne, A Study Course in Organic Chemistry N. Sue Barnes, 1979

Study Guide and Solutions Manual to Accompany Fundamentals of Organic Chemistry McMurry, Susan McMurry, 1990

Study Guide and Solutions Manual to Accompany Fundamentals of Organic Chemistry John McMurry, Susan McMurry, 1986

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PSI Study Guide for Organic Chemistry by Douglas C. Neckers, Michael P. Doyle E. C. Blossey, 1977

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Barron's Educational Series, Joseph Di Rienzo, John J. Ference, Nicole D. Cornell, Edwin H. Hines, John Swartwood, 2018-05-15

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