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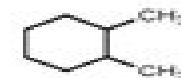
55 multiple-choice questions (3 pts each) + 6 fill-in questions (40 pts) = Total (205 pts).

1) How many unpaired electrons are present in the electron configuration of the O atom?
A) none B) one C) two D) three E) four2) Which of the following contain(s) polar covalent bonds?
A) H₂ B) NaI C) Cl₂ D) H₂O E) both A and C3) What is the formal charge and hybridization for the highlighted Carbon?
A) 0, sp² B) +1, sp² C) -1, sp D) +1, sp E) 0, sp $\text{H}_3\text{C}-\text{N}\equiv\text{C}^+$ 4) What of the following best describes the hybridization of C in carbocation CH₃⁺?
A) sp³ B) sp² C) sp D) s5) what orbital is the nitrogen lone pair residing in the following molecule?
A) sp³ B) sp² C) sp D) p E) s

6) Identify the most acidic carboxylic acid.



Answer is E

7) The conjugate acid of H₂O is
A) HO⁻ B) H₃O C) H₃O⁺ D) H₂O E) H₂O⁺8) How many potential stereoisomers exist for the following molecule?
A) 1 B) 2 C) 3 D) 4 E) 5

9) what is the name of the following compound?

A) *trans*-4-ethylcyclohexan-1-ol B) *cis*-4-ethylhexan-1-ol
C) *trans*-4-methylcyclohexan-1-ol D) *cis*-4-ethylcyclohexan-1-ol
E) *trans*-4-ethylcyclopentan-1-ol

10) which of the following step is part of the propagation in the methane chlorination?

A) $\text{Cl}-\text{Cl} \longrightarrow 2 \text{Cl}^{\cdot}$ B) $\text{H}_3\text{C}^{\cdot} + \text{Cl}-\text{Cl} \longrightarrow \text{H}_3\text{C}-\text{Cl} + \text{Cl}^{\cdot}$
C) $\text{H}_3\text{C}-\text{H} + \text{Cl}^{\cdot} \longrightarrow \text{H}_3\text{C}^{\cdot} + \text{H}-\text{Cl}$ D) B and C

Answer is D

11) Based on the Hammond postulate, which of the following transition state best explains the following reaction step?

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M Walker

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