

# Reactions of Amines and Epoxides

## Preparation of Amines

Amines are often made by reduction of nitriles or by reductive amination of aldehydes and ketones.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

## Bond Types, Strength, Length, Polarity and Dipole Moment

### Two types of covalent bond

There are two types of covalent bond: polar and non-polar covalent bond.

- **Polar covalent bond** is formed by the sharing of electron pairs between atoms of different electronegativity (polar) and non-polar covalent bond.



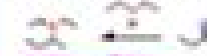
- **Non-polar covalent bond** is formed by sharing of electron pairs between two atoms of the same electronegativity.



### The difference between polar and non-polar covalent bond



### Bond and Electron sharing



### Bonding energy



### There are two types of covalent bond



## Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- **Sigma bonds** are formed by the end-to-end overlap of orbitals.
- **Pi bonds** are formed by the side-to-side overlap of orbitals.
- **Sigma bonds** are stronger than pi bonds.
- **Double bonds** consist of one sigma bond and one pi bond.



pi bonds are weaker than sigma bonds.

## Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be polar or non-polar.



- If the difference in electronegativity is less than 0.5, the bond is non-polar covalent.
- If the difference in electronegativity is between 0.5 and 1.7, the bond is polar covalent.
- If the difference in electronegativity is greater than 1.7, the bond is ionic.

A difference of 1.7 or higher indicates that the electrons are not shared equally and an ionic bond is formed.

A difference of 0.5 or higher indicates that the electrons are not shared equally and a polar covalent bond is formed.

## Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ<sup>-</sup>) and the less electronegative atom has a partial positive charge (δ<sup>+</sup>).
- δ<sup>+</sup> (delta plus) indicates a partial positive charge.
- δ<sup>-</sup> (delta minus) indicates a partial negative charge.

## Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is calculated.

The polarity of the molecule is determined by the vector sum of the dipole moments. If the sum is zero, the molecule is non-polar. If the sum is not zero, the molecule is polar.



Draw the vector sum of the dipole moments to determine the molecular polarity.

## Bond Length and Bond Strength

The length of a bond is the distance between the nuclei of the two atoms.

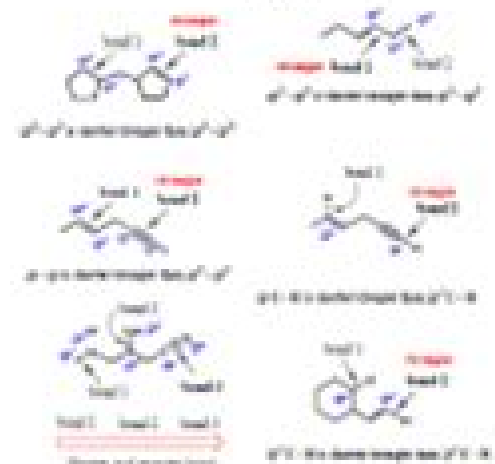


The length of a bond is the distance between the nuclei of the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp <sup>3</sup> -sp <sup>3</sup>	sp <sup>2</sup> -sp <sup>2</sup>	sp-sp
Hybridization	sp <sup>3</sup> -sp <sup>3</sup>	sp <sup>2</sup> -sp <sup>2</sup>	sp-sp
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol
Hybridization	sp <sup>3</sup> -sp <sup>3</sup>	sp <sup>2</sup> -sp <sup>2</sup>	sp-sp
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol

Higher s character means stronger bond

## Examples



# Organic Chem Amine Study Guide

**David R. Klein**



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