

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium chlorides: NH_4^+ + Cl^- \rightarrow $\text{NH}_3 + \text{H}_2\text{O}$ + NH_4Cl



Four different types of bonds

1. Ionic bond
2. Covalent bond
3. Hydrogen bond
4. Metal bond



The Behavior of the four bonds



Four and Five bonds



Four and Five bonds

Four and Five bonds



A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Protophilic Agents



Preparation and Properties

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).

1. 100% ionic bond: $\text{Na}^+ - \text{Cl}^-$
2. 50% ionic bond: $\text{H}^+ - \text{F}^-$
3. 0% ionic bond: $\text{H}^+ - \text{H}^-$



- The dipole moment is the net charge passing from the less electronegative bond to the more electronegative atom.
- The more the dipole moment, the partial negative charge and the less the positive charge.
- It is calculated, relative to the charge difference.
- It is also used to denote the electron density.

Molecular Dipole

To measure the molecular polarity, the total sum of the dipole is used.

The polarity of the molecule is equal to the sum of individual bond dipole moments. If dipole moments are equal, it is just simple addition.



Sum of bond dipole moment = $\mu_{\text{mole}} = \mu_1 + \mu_2 + \mu_3 + \dots$

Bond Length and Bond Strength

Relative bond length



Length: $\text{H}-\text{F} < \text{H}-\text{Cl} < \text{H}-\text{Br} < \text{H}-\text{I}$

Strength: $\text{H}-\text{F} > \text{H}-\text{Cl} > \text{H}-\text{Br} > \text{H}-\text{I}$

1. $\text{H}-\text{F}$ is shorter than $\text{H}-\text{Cl}$.

2. $\text{H}-\text{F}$ is stronger than $\text{H}-\text{Cl}$.

3. $\text{H}-\text{F}$ is shorter than $\text{H}-\text{Br}$.

4. $\text{H}-\text{F}$ is stronger than $\text{H}-\text{Br}$.

5. $\text{H}-\text{F}$ is shorter than $\text{H}-\text{I}$.

6. $\text{H}-\text{F}$ is stronger than $\text{H}-\text{I}$.

Relative bond strength

Strength: $\text{H}-\text{F} > \text{H}-\text{Cl} > \text{H}-\text{Br} > \text{H}-\text{I}$

Average values of C-C and C-H bond length and strength

C-C bond: $\text{H}_2\text{C}-\text{CH}_3$, $\text{H}_2\text{C}=\text{CH}_2$, $\text{CH}_3\text{CH}_2\text{CH}_3$

Hydrogen bond: $\text{H}-\text{C}(=\text{O})-\text{H}$, $\text{H}-\text{C}(=\text{O})-\text{H}$, $\text{H}-\text{C}(=\text{O})-\text{H}$

Length: 1.54 \AA , 1.10 \AA , 1.10 \AA

Strength: 100 kJ/mol , 177 kJ/mol , 177 kJ/mol

Hydrogen bond: $\text{H}-\text{C}(=\text{O})-\text{H}$, $\text{H}-\text{C}(=\text{O})-\text{H}$, $\text{H}-\text{C}(=\text{O})-\text{H}$

Hydrogen bond: $\text{H}-\text{C}(=\text{O})-\text{H}$, $\text{H}-\text{C}(=\text{O})-\text{H}$, $\text{H}-\text{C}(=\text{O})-\text{H}$

Length: 1.10 \AA , 1.10 \AA , 1.10 \AA

Strength: 100 kJ/mol , 100 kJ/mol , 100 kJ/mol

Hydrogen bond and strong bond

Example



$\text{H}-\text{C}(=\text{O})-\text{H}$ is the same length as $\text{C}-\text{C}$.

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David R. Klein

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