

CHEMISTRY FORM THREE NOTES

CHEMISTRY OF CARBON

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A: CARBON

Carbon is an element in Group IV (Group 4) of the Periodic table. It has atomic number 6 and electronic configuration $2:4$ and thus has four valence electrons (tetravalent). It does not easily ionize but forms strong covalent bonds with other elements including itself.

(a) Occurrence

Carbon mainly naturally occurs as:

- (i) allotropes of carbon i.e. graphite, diamond and fullerenes.
- (ii) amorphous carbon in coal, peat, charcoal and coke.
- (iii) carbon(IV) oxide gas accounting 0.03% by volume of normal air in the atmosphere.

(b) Allotropes of Carbon

Carbon naturally occurs in two main crystalline allotropic forms, carbon-graphite and carbon-diamond

Carbon-diamond	Carbon-graphite
Shiny crystalline solid	Black/dull crystalline solid
Has a very high melting/boiling point because it has a very closely packed giant tetrahedral structure joined by strong covalent bonds	Has a high melting/boiling point because it has a very closely packed giant hexagonal planar structure joined by strong covalent bonds
Has very high density (Hardest known natural substance)	Soft
Abrasive	Slippery
Poor electrical conductor because it has no free delocalized electrons	Good electrical conductor because it has free 4^{th} valency delocalized electrons
Is used in making Jewels, drilling and cutting metals	Used in making Lead-pencils, electrodes in batteries and as a lubricant
Has giant tetrahedral structure	Has giant hexagonal planar structure

c) Properties of Carbon

(i) Physical properties of carbon

Carbon occurs widely and naturally as a black solid

It is insoluble in water but soluble in carbon disulphide and organic solvents.

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K Morrison



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Table of Contents Pdf Chemistry Form Three

1. Understanding the eBook Pdf Chemistry Form Three
 - The Rise of Digital Reading Pdf Chemistry Form Three
 - Advantages of eBooks Over Traditional Books
2. Identifying Pdf Chemistry Form Three
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pdf Chemistry Form Three
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pdf Chemistry Form Three

- Personalized Recommendations
- Pdf Chemistry Form Three User Reviews and Ratings
- Pdf Chemistry Form Three and Bestseller Lists
- 5. Accessing Pdf Chemistry Form Three Free and Paid eBooks
 - Pdf Chemistry Form Three Public Domain eBooks
 - Pdf Chemistry Form Three eBook Subscription Services
 - Pdf Chemistry Form Three Budget-Friendly Options
- 6. Navigating Pdf Chemistry Form Three eBook Formats
 - ePub, PDF, MOBI, and More
 - Pdf Chemistry Form Three Compatibility with Devices
 - Pdf Chemistry Form Three Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pdf Chemistry Form Three
 - Highlighting and Note-Taking Pdf Chemistry Form Three
 - Interactive Elements Pdf Chemistry Form Three
- 8. Staying Engaged with Pdf Chemistry Form Three
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pdf Chemistry Form Three
- 9. Balancing eBooks and Physical Books Pdf Chemistry Form Three
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pdf Chemistry Form Three
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pdf Chemistry Form Three
 - Setting Reading Goals Pdf Chemistry Form Three
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pdf Chemistry Form Three

- Fact-Checking eBook Content of Pdf Chemistry Form Three
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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