

- **Thermodynamics** is the study of energy and its transformations.
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.

$$\Delta U = Q - W$$

 where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.
 - **Second Law of Thermodynamics**: The entropy of an isolated system always increases over time.

$$\Delta S \geq 0$$

 where ΔS is the change in entropy.



- **Enthalpy** (H) is a thermodynamic property that represents the total energy of a system, including internal energy and the energy required to displace the environment.

$$H = U + PV$$

 where U is internal energy, P is pressure, and V is volume.
 - **Enthalpy Change** (ΔH) is the change in enthalpy during a process.

$$\Delta H = H_{\text{final}} - H_{\text{initial}}$$

 - **Exothermic Process**: A process that releases heat to the surroundings, resulting in a negative ΔH .
 - **Endothermic Process**: A process that absorbs heat from the surroundings, resulting in a positive ΔH .



- **Free Energy** (G) is a thermodynamic property that represents the maximum useful work that can be extracted from a system at constant temperature and pressure.

$$G = H - TS$$

 where H is enthalpy, T is temperature, and S is entropy.
 - **Gibbs Free Energy Change** (ΔG) is the change in Gibbs free energy during a process.

$$\Delta G = \Delta H - T\Delta S$$

 - **Spontaneous Process**: A process that occurs without the need for external energy input, resulting in a negative ΔG .
 - **Non-spontaneous Process**: A process that requires external energy input to occur, resulting in a positive ΔG .



- **Chemical Equilibrium** is a state in which the concentrations of reactants and products remain constant over time.

$$K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

 where K is the equilibrium constant, $[A]$, $[B]$, $[C]$, and $[D]$ are the concentrations of the reactants and products, and a , b , c , and d are the stoichiometric coefficients.
 - **Le Chatelier's Principle**: If a system at equilibrium is subjected to a change in concentration, temperature, or pressure, the system will shift in a direction that counteracts the change.



- **Electrochemistry** is the study of the relationship between electrical energy and chemical reactions.
 - **Redox Reaction**: A reaction in which one species is oxidized (loses electrons) and another species is reduced (gains electrons).

$$\text{Oxidation: } \text{M} \rightarrow \text{M}^+ + e^-$$

$$\text{Reduction: } \text{M}^+ + e^- \rightarrow \text{M}$$

 - **Electrode Potential** (E°) is a measure of the tendency of a species to be oxidized or reduced.

$$E^\circ_{\text{cell}} = E^\circ_{\text{cathode}} - E^\circ_{\text{anode}}$$

 - **Nernst Equation**: A equation that relates the electrode potential to the concentrations of the reactants and products.

$$E = E^\circ - \frac{RT}{nF} \ln Q$$

 where E is the electrode potential, E° is the standard electrode potential, R is the gas constant, T is the temperature, n is the number of electrons transferred, F is the Faraday constant, and Q is the reaction quotient.



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David C. Colander



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