

Chemistry Study Guide

Periodic Table of the Elements																		
H	Li	Be																
Na	Mg																	
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr	
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	
Cs	Ba	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn		
Fr	Ra	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Uut	Fl	Uup	Lv	Uus	Uuo		
		La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu		
		Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr		

Objectives:

To understand the structure of an atom, and the nature + properties of subatomic particles.



Matter:

Anything that has mass and occupies space.

Examples: air, gases, liquids, solidz!

Atoms:

Atoms are the smallest, indivisible particle.

- Outer part → shell/orbit/energy level → carries the electronz (e-)
- Inner part → nucleus → houses the protonz (p+) and neutronz (n)
- Outer & inner part → together these are the subatomic particles.

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p ⁺	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e ⁻	-0	-1	Orbit/Shell

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Carlo Cohen

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Table of Contents Odysseyware Chemistry Study Guide

1. Understanding the eBook Odysseyware Chemistry Study Guide
 - The Rise of Digital Reading Odysseyware Chemistry Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Odysseyware Chemistry Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Odysseyware Chemistry Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Odysseyware Chemistry Study Guide
 - Personalized Recommendations
 - Odysseyware Chemistry Study Guide User Reviews and Ratings
 - Odysseyware Chemistry Study Guide and Bestseller Lists
5. Accessing Odysseyware Chemistry Study Guide Free and Paid eBooks
 - Odysseyware Chemistry Study Guide Public Domain eBooks

- Odysseyware Chemistry Study Guide eBook Subscription Services
- Odysseyware Chemistry Study Guide Budget-Friendly Options

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- ePub, PDF, MOBI, and More
- Odysseyware Chemistry Study Guide Compatibility with Devices
- Odysseyware Chemistry Study Guide Enhanced eBook Features

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- Adjustable Fonts and Text Sizes of Odysseyware Chemistry Study Guide
- Highlighting and Note-Taking Odysseyware Chemistry Study Guide
- Interactive Elements Odysseyware Chemistry Study Guide

8. Staying Engaged with Odysseyware Chemistry Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Odysseyware Chemistry Study Guide

9. Balancing eBooks and Physical Books Odysseyware Chemistry Study Guide

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Odysseyware Chemistry Study Guide

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Odysseyware Chemistry Study Guide

- Setting Reading Goals Odysseyware Chemistry Study Guide
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Odysseyware Chemistry Study Guide

- Fact-Checking eBook Content of Odysseyware Chemistry Study Guide
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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web 1 obtain a sample of beanium separate the pintonium navium and lentilium and then mass each group record the values in the data table 2 count the number of each isotope by counting the pintonium navium and lentilium 3 calculate the total number of isotopes

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web procedure obtain a sample of the classroom mixture of beanium sort your beanium into its three isotopic bean types count the number of beans in each pile find the mass of each pile of beans determine the average mass of each type of bean based on the samples masses

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web good luck saftey 1 do not eat the beans 2 do not throw the beans 3 do not crush the beans don t mess with the beans pre lab questions 1 there are many beans in your container each bean represents an 2 there are three kinds of beans in your container each kind represents a different