

# Chemistry Study Guide

# Periodic Table of the Elements

<small>1</small> <b>H</b> <small>1.008</small>																	<small>2</small> <b>He</b> <small>4.003</small>																															
<small>3</small> <b>Li</b> <small>6.941</small>	<small>4</small> <b>Be</b> <small>9.012</small>																	<small>5</small> <b>B</b> <small>10.811</small>	<small>6</small> <b>C</b> <small>12.011</small>	<small>7</small> <b>N</b> <small>14.007</small>	<small>8</small> <b>O</b> <small>15.999</small>	<small>9</small> <b>F</b> <small>18.998</small>	<small>10</small> <b>Ne</b> <small>20.180</small>																									
<small>11</small> <b>Na</b> <small>22.990</small>	<small>12</small> <b>Mg</b> <small>24.305</small>																	<small>13</small> <b>Al</b> <small>26.982</small>	<small>14</small> <b>Si</b> <small>28.086</small>	<small>15</small> <b>P</b> <small>30.974</small>	<small>16</small> <b>S</b> <small>32.065</small>	<small>17</small> <b>Cl</b> <small>35.453</small>	<small>18</small> <b>Ar</b> <small>39.948</small>																									
<small>19</small> <b>K</b> <small>39.098</small>	<small>20</small> <b>Ca</b> <small>40.078</small>	<small>21</small> <b>Sc</b> <small>44.956</small>	<small>22</small> <b>Ti</b> <small>47.883</small>	<small>23</small> <b>V</b> <small>50.942</small>	<small>24</small> <b>Cr</b> <small>51.996</small>	<small>25</small> <b>Mn</b> <small>54.938</small>	<small>26</small> <b>Fe</b> <small>55.845</small>	<small>27</small> <b>Co</b> <small>58.933</small>	<small>28</small> <b>Ni</b> <small>58.693</small>	<small>29</small> <b>Cu</b> <small>63.546</small>	<small>30</small> <b>Zn</b> <small>65.38</small>	<small>31</small> <b>Ga</b> <small>69.723</small>	<small>32</small> <b>Ge</b> <small>72.630</small>	<small>33</small> <b>As</b> <small>74.922</small>	<small>34</small> <b>Se</b> <small>78.96</small>	<small>35</small> <b>Br</b> <small>79.904</small>	<small>36</small> <b>Kr</b> <small>83.80</small>																															
<small>37</small> <b>Rb</b> <small>85.468</small>	<small>38</small> <b>Sr</b> <small>87.62</small>	<small>39</small> <b>Y</b> <small>88.906</small>	<small>40</small> <b>Zr</b> <small>91.224</small>	<small>41</small> <b>Nb</b> <small>92.906</small>	<small>42</small> <b>Mo</b> <small>95.94</small>	<small>43</small> <b>Tc</b> <small>98.906</small>	<small>44</small> <b>Ru</b> <small>101.07</small>	<small>45</small> <b>Rh</b> <small>102.91</small>	<small>46</small> <b>Pd</b> <small>106.42</small>	<small>47</small> <b>Ag</b> <small>107.87</small>	<small>48</small> <b>Cd</b> <small>112.41</small>	<small>49</small> <b>In</b> <small>114.82</small>	<small>50</small> <b>Sn</b> <small>118.71</small>	<small>51</small> <b>Sb</b> <small>121.76</small>	<small>52</small> <b>Te</b> <small>127.6</small>	<small>53</small> <b>I</b> <small>126.91</small>	<small>54</small> <b>Xe</b> <small>131.29</small>																															
<small>55</small> <b>Cs</b> <small>132.91</small>	<small>56</small> <b>Ba</b> <small>137.33</small>			<small>57</small> <b>Hf</b> <small>178.49</small>	<small>58</small> <b>Ta</b> <small>180.95</small>	<small>59</small> <b>W</b> <small>183.84</small>	<small>60</small> <b>Re</b> <small>186.21</small>	<small>61</small> <b>Os</b> <small>190.23</small>	<small>62</small> <b>Ir</b> <small>192.22</small>	<small>63</small> <b>Pt</b> <small>195.08</small>	<small>64</small> <b>Au</b> <small>196.97</small>	<small>65</small> <b>Hg</b> <small>200.59</small>	<small>66</small> <b>Tl</b> <small>204.38</small>	<small>67</small> <b>Pb</b> <small>207.2</small>	<small>68</small> <b>Bi</b> <small>208.98</small>	<small>69</small> <b>Po</b> <small>209</small>	<small>70</small> <b>At</b> <small>210</small>	<small>71</small> <b>Rn</b> <small>222</small>																														
<small>87</small> <b>Fr</b> <small>223</small>	<small>88</small> <b>Ra</b> <small>226</small>			<small>89</small> <b>Rf</b> <small>261</small>	<small>90</small> <b>Db</b> <small>262</small>	<small>91</small> <b>Sg</b> <small>266</small>	<small>92</small> <b>Bh</b> <small>264</small>	<small>93</small> <b>Hs</b> <small>277</small>	<small>94</small> <b>Mt</b> <small>268</small>	<small>95</small> <b>Ds</b> <small>271</small>	<small>96</small> <b>Rg</b> <small>272</small>	<small>97</small> <b>Cn</b> <small>285</small>	<small>98</small> <b>Uut</b> <small>289</small>	<small>99</small> <b>Fl</b> <small>288</small>	<small>100</small> <b>Uup</b> <small>289</small>	<small>101</small> <b>Lv</b> <small>293</small>	<small>102</small> <b>Uus</b> <small>294</small>	<small>103</small> <b>Uuo</b> <small>294</small>																														
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## Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

## Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

## Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons ( $e^-$ )
- **Inner part** → *nucleus* → houses the protons ( $p^+$ ) and neutrons ( $n$ )
- Outer & Inner part → together these are the **subatomic particles**

## Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	$p^+$	1	+1	Nucleus
Neutron	$n$	1	0	Nucleus
Electron	$e^-$	$\sim 0$	-1	Orbit/Shell

# Odysseyware Chemistry Study Guide

**Carlo Cohen**



## Odysseyware Chemistry Study Guide:

*Chemistry for Changing Times* John W. Hill, 1988      **Study Guide for Chemistry** Jill Kirsten Robinson, John E. McMurry, Robert C. Fay, 2019-01-04      *The Key Student Study Guide*, 2023      **Study Guide for Chemistry H** LeMay, Theodore Brown, Bruce Bursten, Catherine Murphy, Patrick Woodward, Matthew Stoltzfus, 2017-03-21 This guide assists students through the text material with chapter overviews learning objectives a review of key terms as well as self tests with answers and explanations This student guide also features MCAT practice questions      **Student's Study Guide and Selected Solution Manual for Chemistry for Changing Times** John W. Hill, Terry W. McCreary, Doris K. Kolb, Richard Jones, Marcia Gillette, 2015-04-02 The Study Guide and Selected Solutions Manual assists students with the text material It contains learning objectives chapter outlines additional problems with self tests and answers and answers to the odd numbered problems in the text      **Study Guide and Selected Solutions Manual for Chemistry for Changing Times** John W. Hill, Richard Jones, Doris K. Kolb, Terry W. McCreary, Marcia L. Gillette, 2012-01-31      **General Chemistry II, Student Study Guide** Enrique Olivas, 2015-08-26      Introduction to Chemistry Arthur L. Williams, 1981      Oxford Resources for IB DP Chemistry Geoffrey Neuss, 2023-10-23 Please note this title is suitable for any student studying Exam Board International Baccalaureate IB Level and subject Diploma Programme DP Chemistry First teaching 2023 First exams 2025 The Oxford Resources for IB DP Chemistry Study Guide is an accessible student friendly resource fully aligned to and focused on the knowledge contents of the 2023 DP Chemistry subject guide It is designed to be used alongside the Course Book to help students focus on crucial concepts and skills to build confidence reinforce essential theory and cement understanding of SL and HL ideas in an easy to digest bitesize format Concise explanations diagrams and practical notes engage learners and provide a supportive framework for developing subject comprehension and encouraging a good approach to revision Clear and accessible language throughout supports EAL learners      Chemistry Study Guide and Student Solutions Guide Sixth Edition Zumdahl, 2003-01-01      **Chemistry** Debbie Finocchio, 2006      **Chemistry Lesson Plans, Study Guides, and Notes** Doctor Chemistry, 2017-08-11 The Ultimate Guide to Learning or Teaching Chemistry This book contains the real lecture notes and slide of a highly effective high school and college Chemistry teacher Teachers Never plan another lesson again Students Ace your upcoming exam This series covers all of the topics of High School Chemistry and General Chemistry including Accuracy and Significant Figures Mixtures Metric System Bonding Geometric Bonding Stoichiometry States of Matter Gas Laws and Calculations Reaction Calculations Limiting Reagents Electro Chemistry Organic Chemistry Basics      *The World of Chemistry Study Guide* Gilbert William Castellan, 1990      *Study Guide and Selected Solutions Manual for Basic Chemistry* Karen C. Timberlake, William Timberlake, 2019-06-06 The Study Guide and Selected Solutions Manual as written specifically to assist students using Basic Chemistry It contains learning objectives chapter outlines additional problems with self tests and answers and answers to the odd numbered problems in the text

**Study Guide for Chemistry** Nivaldo J. Tro, Jennifer Shanoski, 2017-01-27 This Study Guide was written specifically to assist students using Structure and Properties It presents the major concepts theories and applications discussed in the text in a comprehensive and accessible manner for students It contains learning objectives chapter summaries and outlines as well as examples self tests and concept questions Fundamentals of Chemistry James E. Brady, 1984 **Complete**

**Chemistry Study Guide + Flashcards and MCAT Review** Konstantinos Papadopoulos, 2013-05-14 This book is part of The Dean's Guide series including study guides on biology chemistry physics and organic chemistry A portion of the proceeds from each book will go to charity sponsoring a less fortunate child The author of the Dean's Guide to Chemistry is a student himself who has chosen to write a simplified version of chemistry that can be used for chemistry at any level high school AP chemistry college and even the MCAT The author has published scientific publications on malaria and is currently pursuing his M S at the University of Glasgow in Infection Immunobiology researching under the director of the institute on rheumatoid arthritis The author uses simple real life examples and illustrations to allow students at any level to understand chemistry The study guide is based on a simplified method which removes all the unnecessary wording and terminology you find in your typical textbook lecture and replaces it with easy to understand relatable examples and wording We all know how hard chemistry can be especially when you have to sit in auditorium of 300 students listening to someone who likes like they're the size of an ant So why not purchase a simplified guide to chemistry and learn chemistry easier Understand these concepts in much less time than normal and go get that A The author is a self-described science enthusiast who believes the real problem with science today is the complexity of the textbooks and teaching methods given to students As a student himself he has realized this through numerous high level science courses and found out he could create a simplified study aide to chemistry that explains some of the hardest topics in very easy ways What the author has done is thrown out all the unnecessary vocabulary equations etc and created a book that is filled with exactly what you should know regarding chemistry through easy to understand examples and step by step solutions with the author's personal explanations

*Chemistry Lesson Plans, Study Guides, and Lecture Notes* Doctor Chemistry, 2017-08-13 The Ultimate Guide to Learning or Teaching Chemistry This book contains the real lecture notes and slides of a highly effective high school and college Chemistry teacher Teachers Never plan another lesson again Students Ace your upcoming exam This series covers all of the topics of High School Chemistry and General Chemistry including Accuracy and Significant Figures Mixtures Metric System Bonding Atomic Theory Periodic Table VSEPR Ionic and Covalent Bonding Geometric Bonding The Mole and Molar Mass Equation Balancing Thermodynamics Stoichiometry States of Matter Gas Laws and Calculations Reaction Calculations Acids and Bases Limiting Reagents Redox and Electro Chemistry Organic Chemistry Basics **Help-u-science** Carlo Cohen, 1999 Chemistry Paul B. Kelter, 1989

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