

Study Guide: Thermal Energy

I. What is Thermal Energy?

- **Definition:** The total kinetic and potential energy of all the particles in a substance.
- **Relationship to Particle Motion:** The faster the particles move, the more thermal energy they have.
- **Factors Affecting Thermal Energy:**
 - Temperature
 - Mass
 - State of matter (solid, liquid, gas)

II. Temperature

- **Definition:** A measure of the *average* kinetic energy of the particles in a substance.
- **Units:** Degrees Celsius ($^{\circ}\text{C}$), degrees Fahrenheit ($^{\circ}\text{F}$), Kelvin (K)
- **Relationship to Thermal Energy:** Higher temperature means higher average kinetic energy, which generally means more thermal energy.

III. Heat

- **Definition:** The transfer of thermal energy from a warmer object to a cooler object.
- **Direction of Heat Flow:** Heat always flows from hot to cold until thermal equilibrium is reached.
- **Units:** joules (J), calories (cal)
- **Methods of Heat Transfer:**
 - **Conduction:** Heat transfer through direct contact.
 - Example: Touching a hot stove.
 - **Convection:** Heat transfer through the movement of fluids (liquids or gases).
 - Example: Boiling water in a pot.
 - **Radiation:** Heat transfer through electromagnetic waves.
 - Example: Feeling the heat from the sun.

IV. Thermal Properties of Matter

- **Thermal Expansion and Contraction:** Most substances expand when heated and contract when cooled.
 - **Explanation:** Particles move farther apart when heated and closer together when cooled.

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M Lipman



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