
ACS ANALYTICAL CHEMISTRY STUDY GUIDE

- K_f : $K_f = K_f \cdot a$
 - EDTA titrations: 1. before eq EDTA binds to metal 1:1
2. At eq we assume that all metal ions can bind to EDTA and then use K_f
- Potentiometry: an analytical method in which an electric potential difference (avoltage) of a cell is measured
-
- Reducing agent: Donates electrons and becomes oxidized.
 - oxidizing agent: Accepts electrons and becomes reduced.
 - coulometry: science of measuring change and relating it to analyte
 - cathode: the electrode at which reduction occurs
 - Anode: the electrode at which oxidation occurs
 - E_{cell} : $E_{\text{cell}} = E_{\text{red}} - E_{\text{ox}}$
 - Faraday's Equation: $E = E - (.0592/n) \log(a/b)$
 - ||: anode always written on left of ||
 - amperometry: elec techniques where the cell current is measured where pot diff between the indicator and reference electrode is controlled
 - Clark electrode: Electrode for oxygen
 - Biosensors: Detecting devices that can detect specific substances in cells
 - amp chromo det: point, size, or charge
 - limiting current: the diff between current and steady state max
 - Anodic vol: Very sensitive for analyzing metals into ppb
 - fluorescence: used to measure singlets, atoms like rigid aromatic rings, high sensitivity
 - phosphorescence: involves emission of light trip- singlet, much lower than fluoro
 - Transmittance: $A = -\log(T)$ or $2 - \log(T\%)$
 - atomic absorption: lamp-flame-mono-detector
 - atomic emission: flame-mono-detector
 - AA hollow cathode lamp: used to form very narrow band lights
 - AA atomizer: Heat sample to very high temps to break all chemical bonds, hot enough to

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that concept The final section is a series of practice problems to test the concepts collectively The key is provided on a separate page for all study and practice problems **National Union Catalog** ,1979 Includes entries for maps and atlases

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