

| Topic                | Grade | What do you need to know   |  or  | What do You have to work on? |
|----------------------|-------|--|--|------------------------------|
| Atomic Structure     | A     | <ul style="list-style-type: none"> <li>Why an atom is neutral in terms of sub-atomic particles</li> <li>Outline isotopes from data about numbers of protons and neutrons</li> </ul>  |  |                              |
|                      | B     | <ul style="list-style-type: none"> <li>Know the charges and masses of protons, electrons and neutrons.</li> <li>Know how isotopes are the same but have different mass number.</li> </ul>  |  |                              |
|                      | C     | <ul style="list-style-type: none"> <li>Describe the arrangement of an atoms electron structure - 2.8.8.2</li> <li>Describe the arrangement of elements in the periodic table.</li> </ul>   |  |                              |
| Ionic Bonding        | A     | <ul style="list-style-type: none"> <li>Explain ionic bonding using the 'dot and cross' model.</li> <li>Explain the melting point of NaCl by the structure of and bonding.</li> <li>Explain why the melting point of NaCl is lower than that of MgO.</li> </ul>   |  |                              |
|                      | B     | <ul style="list-style-type: none"> <li>Predict the properties of substances with giant ionic structures.</li> <li>Understand that atoms with a full outer shell are electronically stable.</li> </ul>  |  |                              |
|                      | C     | <ul style="list-style-type: none"> <li>Ionic bonds are formed from metals and non-metals</li> <li>Electrons are donated and received.</li> <li>Metals always donate the electron(s).</li> <li>Non- metals always receive the electron(s)</li> </ul>  |  |                              |
| The Periodic Table   | A     | <ul style="list-style-type: none"> <li>Explain some covalent bonds using a 'dot and cross' model</li> <li>Explain double covalent binds using a 'dot and cross' model</li> </ul>   |  |                              |
|                      | B     | <ul style="list-style-type: none"> <li>Know why CO<sub>2</sub> and H<sub>2</sub>O have low melting points</li> </ul>   |  |                              |
|                      | C     | <ul style="list-style-type: none"> <li>Covalent bonds always have non-metals and share electrons</li> <li>Know that the group number is the same as the number of electrons in the outer shell.</li> <li>Be able to find out how many shells an atom has from the period it is in.</li> </ul>                  |  |                              |
| The Group 1 Elements | A     | <ul style="list-style-type: none"> <li>Construct a balanced symbol equation for Cu<sup>2+</sup>, Fe<sup>2+</sup> and Fe<sup>3+</sup> to make Cu(OH)<sub>2</sub>, Fe(OH)<sub>2</sub>, Fe(OH)<sub>3</sub>.</li> <li>Reactivity with water gives NaOH and KOH.</li> <li>Know the process is oxidation.</li> </ul> |  |                              |
|                      | B     | <ul style="list-style-type: none"> <li>Why group 1 elements form similar positive ions.</li> <li>Construct a balance symbol equation of group 1 with water</li> </ul>  |  |                              |
|                      | C     | <ul style="list-style-type: none"> <li>Metal + Water <math>\longrightarrow</math> Metal Hydroxide + Hydrogen</li> <li>Know how to use a flame test to identify potassium, lithium and sodium compounds.</li> </ul>   |  |                              |
| The Group 7 Elements | A     | <ul style="list-style-type: none"> <li>Predict if displacement reactions will occur with other halogens</li> <li>Explain the trend of reactivity of group 7 by ideas of electron gain</li> <li>Explain why a process is reduced from its ionic</li> </ul>  |  |                              |

# Ocr Gateway Separate Sciences B C4 Checklist

**Dorothy Warren**



## **Ocr Gateway Separate Sciences B C4 Checklist:**

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