

Test Bank of Organic Chemistry 4th Edition By Janice Smith

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Chapter 1: Structure and Bonding

1. What is the ground-state electronic configuration of a carbon atom?
A) $1s^2, 2s^2, 2p^3$ B) $1s^2, 2s^2, 2p^2$ C) $1s^2, 2s^2, 2p^4$ D) $1s^2, 2s^2, 2p^4$
2. What is the ground-state electronic configuration of a fluorine atom?
A) $1s^2, 2s^2, 2p^2$ B) $1s^2, 2s^2, 2p^3$ C) $1s^2, 2s^2, 2p^4$ D) $1s^2, 2s^2, 2p^5$
3. What is the ground-state electronic configuration of a magnesium cation (Mg^{2+})?
A) $1s^2, 2s^2, 2p^6$ C) $1s^2, 2s^2, 2p^6, 3s^2$
B) $1s^2, 2s^2, 2p^6, 3s^1$ D) $1s^2, 2s^2, 2p^6, 3s^2, 3p^2$
4. What is the ground-state electronic configuration of a chlorine anion (Cl^{-})?
A) $1s^2, 2s^2, 2p^6$ C) $1s^2, 2s^2, 2p^6, 3s^2, 3p^2$
B) $1s^2, 2s^2, 2p^6, 3s^2, 3p^6$ D) $1s^2, 2s^2, 2p^6, 3s^2, 3p^4$
5. Which of the following statements about valence electrons is true?
A) They are the most tightly held electrons.
B) They do not participate in chemical reactions.

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DP Hallahan



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