

NCERT Solutions for Class-XI Chemistry

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Chapter-2

NCERT Chemistry Class 11

1. Calculate the molecular mass of the following:
 - (i) Calculate the number of electrons which will together weigh one gram.
 - (ii) Calculate the mass and charge of one mole of electrons.
1. (i) Mass of one electron = 9.10939×10^{-31} kg
 \therefore Number of electrons that weigh 9.10939×10^{-31} kg = 1
 \therefore Number of electrons that will weigh 1 g (1×10^{-3} kg)
$$= \frac{1}{9.10939 \times 10^{-31} \text{ kg}} \times (1 \times 10^{-3} \text{ kg})$$
$$= 0.1098 \times 10^{-3} + 31$$
$$= 0.1098 \times 10^{28}$$
$$= 1.098 \times 10^{27}$$
 - (ii) Mass of one electron = 9.10939×10^{-31} kg
Mass of one mole of electron = $(6.022 \times 10^{23}) \times (9.10939 \times 10^{-31} \text{ kg})$
 $= 5.48 \times 10^{-7} \text{ kg}$
Charge on one electron = 1.6022×10^{-19} coulomb
Charge on one mole of electron = $(1.6022 \times 10^{-19} \text{ C}) (6.022 \times 10^{23})$
 $= 9.65 \times 10^4 \text{ C}$
2. (i) Calculate the total number of electrons present in one mole of methane.
(ii) Find (a) the total number and (b) the total mass of neutrons in 7 mg of ^{14}C . (Assume that mass of a neutron = 1.675×10^{-27} kg).
(iii) Find (a) the total number and (b) the total mass of protons in 34 mg of NH_3 at STP. Will the answer change if the temperature and pressure are changed?
2. No the answer will not vary with the change in temperature and pressure because the number of subatomic particles like protons, neutrons and electrons is fixed for each and every element and it does not vary with temperature and pressure.
3. How many neutrons and protons are there in the following nuclei?
 $^{12}_6\text{C}$, $^{16}_8\text{O}$, $^{24}_{12}\text{Mg}$, $^{56}_{26}\text{Fe}$, $^{88}_{38}\text{Sr}$
3. $^{13}_6\text{C}$: Atomic mass = 13
Atomic number = Number of protons = 6
Number of neutrons = (Atomic mass) – (Atomic number)
 $= 13 - 6 = 7$
 $^{16}_8\text{O}$:
Atomic mass = 16
Atomic number = 8
Number of protons = 8

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 - The Rise of Digital Reading Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
 - Advantages of eBooks Over Traditional Books
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 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
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 - Features to Look for in an Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
 - User-Friendly Interface
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 - Interactive Elements Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com

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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pdf File 11 Class Chemistry 2 Chapter 2 Part Ncerthelp Com
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Carving Out Dedicated Reading Time
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