

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Interactions

There are two types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron affinity).



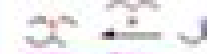
- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Strength of Covalent Bonds



Bond and Bonding Properties



Bonding Properties



There are two types of covalent bonds:

1. Polar covalent bond



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- Sigma bonds are formed by head-on overlap of orbitals.
- Pi bonds are formed by side-on overlap of orbitals.

- A single bond is a sigma bond.
- A double bond consists of one sigma and one pi bond.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity:

• Polar covalent bonds are formed between atoms with different electronegativities.

• Nonpolar covalent bonds are formed between atoms with similar electronegativities.



- If the difference in electronegativity is less than 0.5, the bond is nonpolar covalent.
- If the difference in electronegativity is between 0.5 and 1.7, the bond is polar covalent.
- If the difference in electronegativity is greater than 1.7, the bond is ionic.

- A difference of 1.7 or higher indicates that the electrons are not shared equally and a ionic bond is formed.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is defined as a vector pointing from the less electronegative atom to the more electronegative atom.

- The vector dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

- δ^- (delta minus) indicates an electron deficient atom.

- δ^+ (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is calculated.

The polarity of the molecule is determined by the vector sum of the dipole moments. A polar molecule has a net dipole moment.



Calculate the net dipole moment for the molecule.

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds.

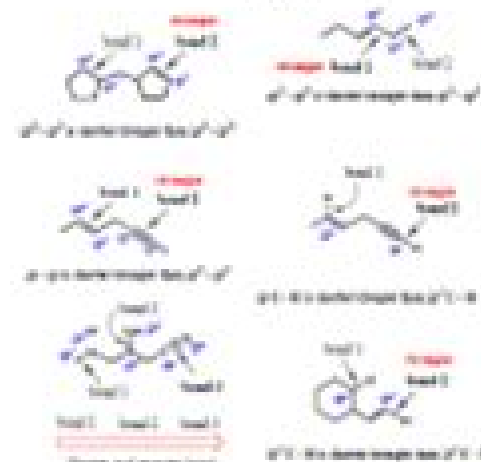


The bond length is the distance between the nuclei of the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Higher electronegativity difference leads to stronger bond.

Examples



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Arthur James Wells



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