

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds



There are four types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electrons from one atom to another (anions and cations) (opposite charges).



Covalent ionic and polar covalent bonds

- Covalent bonds are classified as ionic (polar) and non-polar.
- Ionic bonds are formed by the transfer of electrons.
- Non-polar bonds are formed by the sharing of electrons.
- Polar covalent bonds are formed by the sharing of electrons.
- Non-polar covalent bonds are formed by the sharing of electrons.



Non-polar covalent bonds are stronger than ionic.

Strength of different bonds

Ionic bonds



Covalent bonds



Polar covalent bonds



Non-polar covalent bonds



Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, polar bonds can be polar and nonpolar.

As a rough measure:

- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond. (For example: chlorine and the F, Cl and Br bonds).
- If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond. (Nonpolar covalent bonds are often formed between H and O, N, F and other halogens).
- A difference of 1.7 or higher one of both the electrons are shared and a polar bond is formed.

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Protagonists

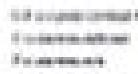


Protagonists and Reactions

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).



Electrons are concentrated around the more electronegative atom.

- The dipole moment is shown as a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector originates at the partial negative end and the arrow points to the partial positive charge end.
- The dipole plot indicates the charge difference.
- The dipole moment indicates the charge distribution.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of the individual bond polarities. If dipole moments are canceled, a polar molecule is called a dipole.



Sum of the bond dipole moment for the molecule.

Bond Length and Bond Strength

Relative bond length



Length: 1.07 Å, 1.27 Å, 1.47 Å, 1.66 Å.

Strength: 0.01 kcal/mol, 0.02 kcal/mol, 0.03 kcal/mol, 0.04 kcal/mol.

Non-polar covalent bond is the weaker and stronger bond.

Polar covalent bond is the weaker and stronger bond.



Relative bond strength

Average values of C-C and C-H bond depending on the substituents

C-C bond: $\text{H}_3\text{C}-\text{CH}_3$, $\text{H}_2\text{C}=\text{CH}_2$, $\text{CH}_3\text{CH}_2\text{CH}_3$

Hybridization: sp^3 (long) sp^2 (medium) sp (short)

Length: 1.70 Å, 1.54 Å, 1.39 Å.

Strength: 0.01 kcal/mol, 0.02 kcal/mol, 0.03 kcal/mol.



Hybridization: sp^3 (long) sp^2 (medium) sp (short)

Length: 1.10 Å, 1.09 Å, 1.08 Å.

Strength: 0.01 kcal/mol, 0.02 kcal/mol, 0.03 kcal/mol.

Hydrogen bond and stronger bond

Example



As a hydrogen bond is formed.



As a hydrogen bond is formed.



As a hydrogen bond is formed.



As a hydrogen bond is formed.

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**Nechamkin, Jacob Sharefkin, George J.
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