

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium chlorides: $\text{NH}_4^+ + \text{H}_3\text{N}^- \rightarrow \text{NH}_3 + \text{NH}_4^+$



Four different types of bonds

1. Ionic bond
2. Covalent bond
3. Hydrogen bond
4. Metallic bond

There are four types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electron from one atom to another (anions and cations). (Opposite charges).



The covalent bonding of two atoms

$\text{O}=\text{C}=\text{O}$

Shared and transferred electrons

H_2

Hydrogen bonding

H_2O

Electrostatic interactions

Na^+

Metals

Mg^{2+}

Hydrogen bonding

H_2O

Hydrogen bonding

H_2O

Hydrogen bonding

H_2O

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Protagonists



Protagonists and Protagonists

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).

1. 100% ionic bond
2. 50% ionic bond
3. 0% ionic bond



Electrostatic map of the oxygen atom.

- The dipole moment is shown as a vector pointing from the less electronegative bond to the more electronegative atom.
- The vector is proportional to the partial negative charge and the distance between a partial positive charge and a partial negative charge.
- The dipole plot indicates the charge difference.
- The dipole moment is shown as a distance and angle.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of the bond dipole moments. If the dipole moments are equal, it is just simple addition.



Sum of the bond dipole moment of the molecule.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, polar bonds can be polar and nonpolar.

- If the difference in electronegativity is less than 0.4, the electrons are almost equally shared between the two atoms. This is a nonpolar covalent bond. (For example: chlorine to the C-C and C-Cl bonds).
- If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond. (Nonpolar covalent bonds are shared by carbon to oxygen in CO_2 , CO and other halogens).
- A difference of 1.7 or higher one of both the electrons are shared by one atom and one bond is formed.

Bond Length and Bond Strength

→ Bond length to length of bond

Bond	$\text{H}-\text{H}$	$\text{H}-\text{O}$	$\text{H}-\text{S}$	$\text{H}-\text{Cl}$
Length	1.07 Å	1.07 Å	1.07 Å	1.07 Å
Strength	100 kJ/mol	117 kJ/mol	120 kJ/mol	127 kJ/mol

Nonpolar covalent bond: same in the nuclei and stronger bond.



→ Bond length to length of bond

Average values of C-C and C-H bond depending on the molecule			
C-C bond	$\text{H}_2\text{C}-\text{CH}_3$	$\text{H}_2\text{C}=\text{CH}_2$	$\text{CH}_3\text{CH}_2\text{CH}_3$
Hybridization	sp^3 hybrid	sp^2 hybrid	sp hybrid
Length	1.54 Å	1.35 Å	1.10 Å
Strength	100 kJ/mol	117 kJ/mol	120 kJ/mol
Hybridization	sp^3	sp^2	sp
Efficiency	25 %	50 %	75 %
Length	1.10 Å	1.07 Å	1.07 Å
Strength	100 kJ/mol	117 kJ/mol	120 kJ/mol

→ Dipole moment and strength bond

Example



$\text{H}-\text{H}$ is a nonpolar bond.

$\text{H}-\text{Cl}$ is a polar bond.

$\text{H}-\text{Cl}$ is a shorter bond.

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AN Whitehead

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