

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium chlorides: $\text{NH}_4^+ + \text{H}_3\text{N}^- \rightarrow \text{NH}_3 + \text{NH}_4^+$



Four different types of bonds

1. Ionic bond
2. Covalent bond
3. Hydrogen bond
4. Metallic bond

There are four types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electron from one atom to another (anions and cations). (Opposite charges).



The covalent bonding of two atoms

$\text{O}=\text{C}=\text{O}$

Shared and transferred electrons

H_2

Hydrogen bonding

H_2O

Electrostatic interactions

Na^+

Metals

Mg^{2+}

Hydrogen bonding

H_2O

Hydrogen bonding

H_2O

Hydrogen bonding

H_2O

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Protagonists



Protagonists and Reactions

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).

1. 100% ionic bond
2. 50% ionic bond
3. 0% ionic bond



Electrostatic map of the molecule showing the charge density.

- The dipole moment is shown as a vector pointing from the less electronegative bond to the more electronegative atom.
- The vector is proportional to the partial negative charge and the electronegativity, even if partial positive charge also.
- The dipole plot indicates the charge difference.
- The dipole moment is shown as a vector with arrows.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of the bond dipole moments. If the dipole moments are equal, it is just simple addition.



Sum of the bond dipole moment of the molecule.

Average values of C-C and C-H bond depending on hybridization			
C-C bond	$\text{sp}_2\text{-sp}_2$	$\text{sp}_3\text{-sp}_3$	$\text{sp}\text{-sp}$
Hybridization	$\text{sp}^2\text{-sp}^2$	$\text{sp}^3\text{-sp}^3$	sp-sp
Length	1.37 Å	1.45 Å	1.55 Å
Strength	100% ionic	50% ionic and 50% covalent	50% ionic and 50% covalent
Hybridization	sp^2	sp^3	sp
Hybridization	sp^2	sp^3	sp
Strength	100% ionic	50% ionic and 50% covalent	50% ionic and 50% covalent
Length	1.37 Å	1.45 Å	1.55 Å
Strength	100% ionic	50% ionic and 50% covalent	50% ionic and 50% covalent

Dipole + bond dipole and stronger bond

Example



$\text{dipole} = \text{dipole}_1 + \text{dipole}_2$

Bond Length and Bond Strength

→ Bond length to length of bond



Bond length = length of bond

→ Bond length to length of bond

Polar and Nonpolar Covalent Bonds

Depends on the difference in the electronegativity

value, polar bond can be polar and nonpolar.

As a rough measure:



If the difference in electronegativity is less than 0.2, the electrons are almost equally shared between the two atoms. This is a nonpolar bond.

If the difference in electronegativity is between 0.3 and 1.7, the electrons are polarized toward the more electronegative atom.

If a difference of 1.8 or higher one or both of the electrons are largely shared with one bond is strong.

→ A difference of 1.8 or higher one or both of the electrons are largely shared with one bond is strong.

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