

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium chlorides: $\text{NH}_4^+ + \text{Cl}^- \rightarrow \text{NH}_3 + \text{HCl}$ (water soluble, soluble in polar solvents)



Four different types of bonds



There are four types of bonding: ionic bonding and covalent bonding

- Ionic bonds are formed by the transfer of electrons from one atom to another (anions and cations) (opposite charges)



Regions (p) and pi (π) bonds

- Covalent bonds are classified as sigma (σ) and pi (π).
- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- Very simple bond is a sigma bond.
- Very complex bond is a pi bond.



pi bonds are stronger than sigma

Dependence of bond length

Electronegativity



Electrostatic Potential



Polar and Nonpolar Covalent Bonds

- Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.
- As a rough measure:

- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond (like a C-C and C-H bond).
- If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond (nonpolar covalent bond electrons are shared between C, O, F and other halogens).
- A difference of 1.7 or higher one or both the electrons are shared with one atom (like C-F).

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Protagonists

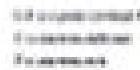


Protagonists and Reactions

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).



$\mu = q \times r$ (where q = charge and r = bond length)

Direction of the dipole moment is from the positive to the negative charge.

- The dipole moment is shown as a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector (dipole moment) starts from a partial negative charge and the vector ends at a partial positive charge.
- The dipole plot indicates the charge difference.
- The dipole moment indicates the charge and area.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of individual bond polarities. The dipole moment is measured. A good example is water, dipole.



Sum of the bond dipole moment for the molecule.

Bond Length and Bond Strength

Relative bond length



Length: C-C > C-O > C-Cl > C-F

Strength: C-C < C-O < C-Cl < C-F

Electronegativity: C < O < Cl < F

Electrostatic Potential: C < O < Cl < F

Electron density: C < O < Cl < F

Electron pair density: C < O < Cl < F

Relative bond strength



Strength: C-C < C-O < C-Cl < C-F

Average values of C-C and C-H bond length and strength

C-C bond: 1.54 \AA 100 kJ/mol 100 kJ/mol

Hydrogen bond: 1.15 \AA 450 kJ/mol 450 kJ/mol

Length: 1.54 \AA 1.15 \AA 1.15 \AA

Strength: 100 kJ/mol 100 kJ/mol 450 kJ/mol

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Organic Chemistry Study Guide For Pcat

**George J. Hademenos, Shaun
Murphree, Kathy A. Zahler, Mark
Whitener, Jennifer M. Warner**

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