

Reactions of Amines and Epoxides

Preparation of Amines

Amines of different hybridization are shown below for hybridization and



A Summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Free electron lone-pairs

There are two types of bonding: **sigma bonding** and **covalent bonding**.

• **Sigma bonds** are formed by the overlap of orbitals from one orbital to another (conspicuous) (positive) and **covalent bonding** (negative) (strong).



• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of σ and π Bonds



Bond and Reaction Sites



Bonding in sp^3 hybridization



There are two types of hybridization

1. **Linear hybridization**



2. **Bent hybridization**



Sigma and pi (π) bonds

Covalent bonds are classified as **sigma (σ)** and **pi (π)**.

• **Sigma bonds** are formed by head-on (end-to-end) overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **One single bond is a sigma bond.**

• **Two double bonds consist of one sigma and one pi bond.**



• **pi bonds are stronger than sigma bonds.**

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be **polar** and **nonpolar**.



• If the difference in electronegativity is **less than 0.5**, the electron is shared equally (nonpolar bond) for two atoms, forming a **nonpolar covalent bond**. (Two important examples are the C-C and C-H bonds).

• If the difference in electronegativity is **between 0.5 and 1.7**, the bond is **polar covalent bond**. (Two important examples are water for carbon a bond to N, O, F and other halogens).

• A difference of **1.7 or higher** is so high that the electron is not longer shared and is **ionic bond** is formed.

Dipole Moment

The polarity of a bond is described by the **dipole moment (μ)**.



• The dipole moment is shown as a **vector pointing from the less electronegative to toward the more electronegative atom**.

• The more electronegative atom has a **partial negative charge (δ^-)** and the less electronegative atom is a **partial positive charge (δ^+)**.

• **δ^- (delta minus) indicates an electron deficient area.**

• **δ^+ (delta plus) indicates an electron rich area.**

Molecular Dipole

To increase the molecular polarity, the **vector sum of the dipole's** is used.

The **polarity of the molecule's** is a **vector sum of the dipole's** (the polarity sum for dipole moment) are combined. A good example is carbon dioxide.



Lower the value is the more **nonpolar** the molecule is.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Substance	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	153.8	133.4	120.6
Strength	376 kJ/mol	473 kJ/mol	581 kJ/mol
Substance	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hydrocarbon	CH_4	C_2H_2	C_2H_2
% s-character	25%	33%	50%
Length	153.8	133.4	120.6
Strength	376 kJ/mol	473 kJ/mol	581 kJ/mol

Higher s-character means shorter bond

Bond Length and Bond Strength

← **Shorter the bond, the stronger it is** →

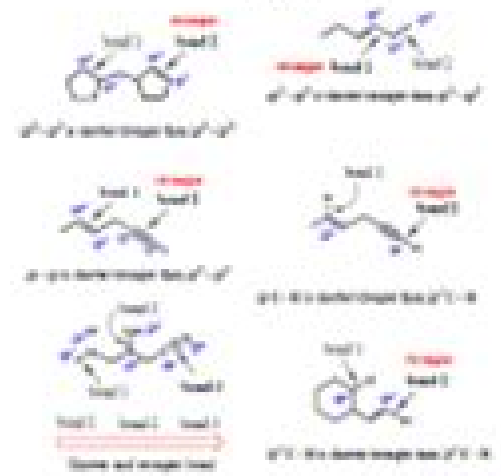
Bond	H-H	H-Cl	H-Br	H-I
Length	74.14	127.4	141.4	154.4
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are **stronger** than single bonds and **shorter** than single bonds.



← **Stronger the bond, the shorter it is** →

Example



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Murphree, Kathy A. Zahler, Mark
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