

Reactions of Amines and Epoxides

Preparation of Amines

Amines are often made by reduction of nitriles or by reductive amination of aldehydes and ketones.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: **sigma bond** and **pi bond**.

- Sigma bonds** are formed by the overlap of atomic orbitals (AOs) in a head-on manner (end-to-end) and are **stronger**.



- Pi bonds** are formed by the overlap of atomic orbitals (AOs) in a side-on manner (side-to-side) and are **weaker**.



The Relative Strength of σ and π Bonds



Bond and Bond Energy

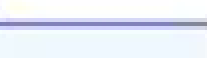


Bond Energy



There are two types of covalent bond

Definition of sigma and pi bonds



Sigma and pi bonds

Covalent bonds are classified into **sigma** and **pi** bonds.

- Sigma bonds** are formed by the overlap of atomic orbitals (AOs) in a head-on manner (end-to-end) and are **stronger**.
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σ bonds are stronger than π bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity of the atoms, covalent bonds can be **polar** or **nonpolar**.



- If the difference in electronegativity is **less than 0.5**, the bond is **nonpolar**.
- If the difference in electronegativity is **between 0.5 and 1.7**, the bond is **polar**.
- If the difference in electronegativity is **greater than 1.7**, the bond is **ionic**.

- A difference of **1.7 or higher** is so high that the electrons are not shared but are **lost** to the more electronegative atom.

- A difference of **1.7 or higher** is so high that the electrons are not shared but are **lost** to the more electronegative atom.

Dipole Moment

The polarity of a bond is described by the **dipole moment** (μ).



- The dipole moment is a vector pointing from the **less electronegative** atom to the **more electronegative** atom.
- The **size** of the dipole moment is a measure of the **polarity** of the bond.
- δ^- (delta minus)** indicates an electron-deficient atom.
- δ^+ (delta plus)** indicates an electron-rich atom.

Molecular Dipole

To determine the **molecular dipole**, the **vector sum** of the **bond dipoles** is used.

The **polarity** of the molecule is a **result** of the **vector sum** of the **bond dipoles**.



Draw the **vector sum** of the **bond dipoles** to find the **molecular dipole**.

Bond Length and Bond Strength

The **length** of a bond is a measure of the **distance** between the two atoms.

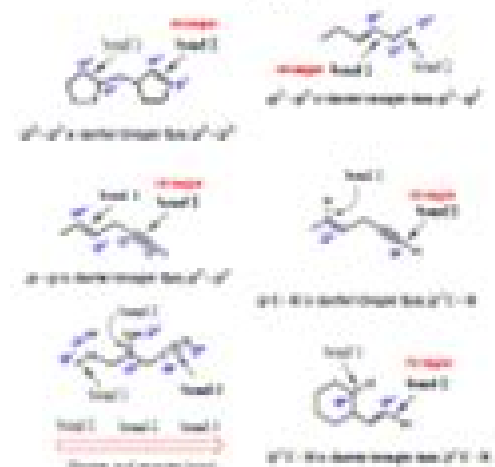


The **length** of a bond is a measure of the **distance** between the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	112 kcal/mol	120 kcal/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.09 Å	1.08 Å	1.06 Å
Strength	99 kcal/mol	110 kcal/mol	120 kcal/mol

Higher s character = stronger and shorter bond

Examples



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