

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Bonding

There are two types of bonding: ionic bonding and covalent bonding.

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron affinity).



- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of Bonds



Bond and Bonding



Bonding



There are two types of bonding

Definition of bonding



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- **Sigma bonds** are formed by head-on overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **Single bonds** are sigma bonds.
- **Double bonds** consist of one sigma and one pi bond.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be polar and nonpolar.



• If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar covalent bond). Examples: C-C and C-H bonds.

• If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar covalent bond). Examples: C-O and C-N bonds.

• A difference of 1.7 or higher indicates an ionic bond (electron transfer).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.

- The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).

- **δ⁻** (delta minus) indicates an electron-deficient atom.

- **δ⁺** (delta plus) indicates an electron-rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the individual bonds are polar. A good example is carbon dioxide.



Cancel the vector as the other pointing point in the same way.

Bond Length and Bond Strength

The longer the bond, the weaker it is.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds and are shorter.

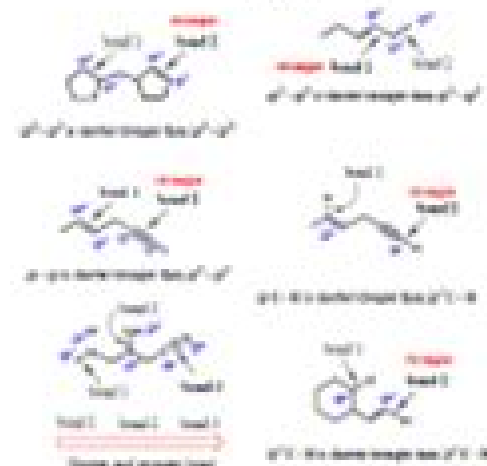


The longer the bond, the weaker it is.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	611 kJ/mol	837 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
C-H Bond	109 pm	108 pm	106 pm
Length	154 pm	134 pm	121 pm
Strength	376 kJ/mol	611 kJ/mol	837 kJ/mol

Higher s character = stronger and shorter bond

Examples



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