

How to Name Chemical Compounds

Is it ionic or covalent?
or does it contain a polyatomic ion followed by H⁺?

Ionic compound

[Cation's name] [Anion's name]

metals

- same as element name
ex. Al³⁺: aluminum
- if transition metal⁺, include charge in Roman numerals
ex. Cu²⁺: copper (II)

poly-
anions

- just like the name!
ex. NH₄⁺: ammonium
- just like H₃O⁺
ex. SO₄²⁻: sulfate

* exceptions (no Roman numerals):

silver ion: silver Ag⁺

gold ion: gold(1+) Au⁺

cadmium ion: cadmium Cd²⁺

Molecular compound

Acids
and
nitrates

[H]-[1st element name] [H]-[2nd element root]-"ide"

↑ skip if "mono-"

↑ skip if 1st element is H

Stem prefixes

- 1: mono- • 6: hexa-
- 2: di- • 7: hepta-
- 3: tri- • 8: octa-
- 4: tetra- • 9: nona-
- 5: penta- • 10: deca-



Acids

Drop all H⁺, and look at the name of the remaining anion.

| Acid | Anion | Acid | Anion |
|--|--|--|--|
| Acid | hydrochloric acid | Acid | Acid |
| ex. HCl ⁺ chloride HCl hydrochloric acid | ex. Ca(OH) ₂ ⁻ oxide Ca(OH) ₂ calcium hydroxide | ex. HCl ⁺ chloride HCl hydrochloric acid | ex. HCl ⁺ chloride HCl hydrochloric acid |
| ex. CH ₃ ⁺ acetate CH ₃ hydrochloric acid | ex. HCl ⁺ chlorate HCl hydrochloric acid | ex. HCl ⁺ chloride HCl hydrochloric acid | ex. HCl ⁺ chloride HCl hydrochloric acid |

Naming Molecular Compounds Study Guide

John McMurry, Susan McMurry

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