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Memes Today Discount

Camilla Rothe



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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation

$$8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$$

$$8.3 \text{ mole } \text{H}_2\text{O} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$$

 149.4 g H_2O is your theoretical yield so percent yield = $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

 if the reaction of 21.3 grams of O_2 produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have

$$\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g } \text{O}_2/\text{mole}} = 0.666 \text{ moles } \text{O}_2$$

$$\frac{21.3 \text{ g } \text{CO}}{28 \text{ g } \text{CO}/\text{mole}} = 0.761 \text{ moles } \text{CO}$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$ in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction

$$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$

 b determine the theoretical yield of Al_2O_3

$$160.0 \text{ g } \text{Al} \times \frac{1 \text{ mol } \text{Al}}{27 \text{ g } \text{Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3/\text{mol} = 150.1 \text{ g } \text{Al}_2\text{O}_3$$

$$302.3 \text{ g } \text{Al}_2\text{O}_3 \times \frac{1 \text{ mol } \text{Al}_2\text{O}_3}{206.0 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 146.6\%$$

 c determine the percent yield

$$\frac{100.260 \text{ g } \text{Al}_2\text{O}_3}{150.1 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 66.8\%$$

 theoretical = 302.3 g

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web 7 Na₂C₂O₄ is the limiting reactant percent yield = 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H_2 are left over 13

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