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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g $\text{Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} = 5.93 \text{ mol Al}$ $\frac{5.93 \text{ mol Al}}{2} = 2.965 \text{ mol Al}_2\text{O}_3$ $2.965 \text{ mol Al}_2\text{O}_3 \times 101.96 \text{ g/mol} = 302.3 \text{ g Al}_2\text{O}_3$ c determine the percent yield $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$ theoretical 302.3 g

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