

# How to Name Chemical Compounds

is it ionic? is an ionic  
or does it contain a polyatomic ion (as provided by H<sup>+</sup>)

## Ionic compound

[cation's name] [anion's name]

main- element	<ul style="list-style-type: none"> <li>take its element name!</li> <li>ex. <math>\text{Al}^{3+}</math>: aluminum</li> <li>if transition metal*, include charge in Roman numeral</li> <li>ex. <math>\text{Cr}^{2+}</math>: chromium (II)</li> </ul>	<ul style="list-style-type: none"> <li>get element's name, chop end, add "-ide"</li> <li>ex. <math>\text{F}^{-}</math>: fluorine → fluoride</li> </ul>
poly- element	<ul style="list-style-type: none"> <li>just use its name!</li> <li>ex. <math>\text{NH}_4^{+}</math>: ammonium</li> </ul>	<ul style="list-style-type: none"> <li>just use its name!</li> <li>ex. <math>\text{CO}_3^{2-}</math>: carbonate</li> </ul>

\* exceptions (no Roman numerals):

silver ion: always  $\text{Ag}^{+}$

zinc ion: always  $\text{Zn}^{2+}$

cadmium ion: always  $\text{Cd}^{2+}$

## Molecular compound

is it  $\text{HClO}_x$ ?

Acid! then use nomenclature

[#]-[1<sup>st</sup> element name] [#]-[2<sup>nd</sup> element root]-"ide"

skip if "mono-" skip if 1<sup>st</sup> element is H

### Numerical prefixes

- 1: mono-
- 2: di-
- 3: tri-
- 4: tetra-
- 5: penta-
- 6: hexa-
- 7: hepta-
- 8: octa-
- 9: nona-
- 10: deca-

for  
first element: never  
ex.  $\text{N}_2\text{O}_5$  → dinitrogen  
pentoxide

### Acids

Ignore all  $\text{H}^{+}$ , and look at the name of the remaining anion.

Anion	X-ide	X-ate	X-ite
Acid	hydro-X-ic acid	X-ic acid	X-ous acid
	ex. $\text{Br}^{-}$ bromide $\text{HBr}$ hydrobromic acid	ex. $\text{C}_2\text{H}_3\text{O}_2^{-}$ acetate $\text{HC}_2\text{H}_3\text{O}_2$ acetic acid	ex. $\text{IO}_3^{-}$ iodate $\text{HIO}_3$ iodic acid
	ex. $\text{CN}^{-}$ cyanide $\text{HCN}$ hydrocyanic acid	ex. $\text{PO}_4^{3-}$ phosphate $\text{H}_3\text{PO}_4$ phosphoric acid	ex. $\text{SO}_3^{2-}$ sulfite $\text{H}_2\text{SO}_3$ sulfurous acid

# Naming Molecular Compounds Study Guide

**Norman E. Griswold, Henry Fuller  
Holtzclaw**



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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement *Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J.

Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter **General**

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Myers, William T. Mooney, 1980 General Chemistry Darrell D. Ebbing, Larry K. Krannich, Joan I. Senyk, 1996 Study

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