

Reactions of Amines and Epoxides

Preparation of Amines

Amines are often made by reduction of nitriles or by reductive amination of aldehydes and ketones.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: ionic and covalent bond.

- Ionic bonds are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron affinity).



- Covalent bonds are formed by sharing of electrons between two atoms.



The difference between ionic and covalent bonds



Bond and Electron sharing



Bonding energy



There are two types of covalent bond



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- Sigma bonds are formed by head-on overlap of orbitals.
- Pi bonds are formed by side-on overlap of orbitals.
- A single bond is a sigma bond.
- A double bond consists of one sigma and one pi bond.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

- A difference of 0.5 or higher can be high for the electrons are not shared (ionic bond).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).
- δ^- (delta minus) indicates an electron deficient atom.
- δ^+ (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments.

The polarity of the molecule is determined by the vector sum of the dipole moments. A good example is carbon dioxide.



Carbon dioxide is a linear molecule with no net dipole moment.

Bond Length and Bond Strength

The length of the bond is measured in Å.

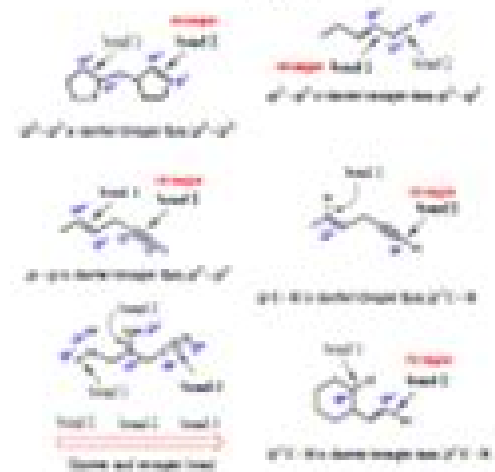


The length of the bond is measured in Å.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Hybridization	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol
Hybridization	sp^3 - sp^3	sp^2 - sp^2	sp - sp
C-H Bond	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Length	1.09 Å	1.06 Å	1.03 Å
Strength	99 kcal/mol	112 kcal/mol	120 kcal/mol

Higher s character = stronger and shorter bond

Example



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Table of Contents Organic Chemistry 3a Study Guide

1. Understanding the eBook Organic Chemistry 3a Study Guide
 - The Rise of Digital Reading Organic Chemistry 3a Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chemistry 3a Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chemistry 3a Study Guide
 - User-Friendly Interface

4. Exploring eBook Recommendations from Organic Chemistry 3a Study Guide
 - Personalized Recommendations
 - Organic Chemistry 3a Study Guide User Reviews and Ratings
 - Organic Chemistry 3a Study Guide and Bestseller Lists
5. Accessing Organic Chemistry 3a Study Guide Free and Paid eBooks
 - Organic Chemistry 3a Study Guide Public Domain eBooks
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 - Organic Chemistry 3a Study Guide Budget-Friendly Options
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 - Highlighting and Note-Taking Organic Chemistry 3a Study Guide
 - Interactive Elements Organic Chemistry 3a Study Guide
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Organic Chemistry 3a Study Guide
9. Balancing eBooks and Physical Books Organic Chemistry 3a Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Organic Chemistry 3a Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Organic Chemistry 3a Study Guide
 - Setting Reading Goals Organic Chemistry 3a Study Guide
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Organic Chemistry 3a Study Guide
 - Fact-Checking eBook Content of Organic Chemistry 3a Study Guide
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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