

Thermochemistry

Multiple Choice Questions

1. When the three types of energy changes are arranged in order of decreasing molar heat values, what is the correct order?
- A. chemical, phase, nuclear
B. nuclear, chemical, phase
C. nuclear, phase, chemical
D. phase, chemical, nuclear
2. The temperature remains constant as energy is added to a substance. How may the substance be changing?
- A. from a gas to a solid
B. from a liquid to a gas
C. from a liquid to a solid
D. in the amount of kinetic energy
3. Which metal requires the most energy to raise 1.00 g of it by 1.00°C?
- A. aluminum
B. copper
C. lead
D. nickel
4. A substance has a molar heat of combustion of -810.4 kJ/mol. When 0.285 mol of the substance is burned in a calorimeter containing 8.60 kg of water, what is the increase of the water temperature?
- A. 0.156°C
B. 6.41°C
C. 7.89°C
D. 12.8°C
5. The addition of 9.54 kJ of heat is required to raise the temperature of 225.0 g of a liquid hydrocarbon from 20.5°C to 45.0°C. What is the heat capacity of this hydrocarbon?
- A. 0.94 J/g°C
B. 1.73 J/g°C
C. 1.88 J/g°C
D. 9.42 J/g°C
6. If the molar heat of fusion of sodium is 2.63 kJ/mol, how much energy is needed to melt 180.0 g of solid sodium at its melting point?
- A. 2.63 kJ
B. 2.93 kJ
C. 28.6 kJ
D. 473 kJ
7. Which process is exothermic?

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Elizabeth Kean, Catherine Middlecamp



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web the real power absorbed by the load is given by $p = 120 \cdot 6.63 \cos 56.4^\circ = 440 \text{ W}$ which can be checked by $i^2 R = 6.63^2 \cdot 10 = 440 \text{ W}$

the reactive power absorbed by the load is $q = 120.663 \sin 36.4^\circ = 663 \text{ var}$ c peak magnetic energy $w_{\text{li}} = \frac{1}{2} L I^2 = \frac{1}{2} (0.04) (6.63)^2 = 0.877 \text{ J}$ ω is satisfied 2.17 A $\text{svi} = \text{zii} = \text{zli} = 2.2 \text{ } \omega$

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