

# Science

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(Chapter – 1) (Matter in Our Surroundings)

(Class – IX)

## Question 4:

Liquids generally have lower density as compared to solids. But you must have observed that ice floats on water. Find out why.

## Answer 4:

The mass per unit volume of a substance is called density (density = mass/volume).

*As the volume of a substance increases, its density decreases.*

Though ice is a solid, it has large number of empty spaces between its particles. These spaces are larger as compared to the spaces present between the particles of water. Thus, the volume of ice is greater than that of water. Hence, the density of ice is less than that of water. A substance with lower density than water can float on water. Therefore, ice floats on water.

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## Question 1:

Convert the following temperatures to the celsius scale.

(a) 300 K

(b) 573 K

## Answer 1:

(a)  $300\text{ K} = (300 - 273)^{\circ}\text{C} = 27^{\circ}\text{C}$

(b)  $573\text{ K} = (573 - 273)^{\circ}\text{C} = 300^{\circ}\text{C}$

## Question 2:

Convert the following temperatures to the celsius scale.

(a)  $250^{\circ}$

(b)  $100^{\circ}$

## Answer 2:

(a) Water at  $250^{\circ}\text{C}$  exists in gaseous state.

(b) At  $100^{\circ}\text{C}$ , water can exist in both liquid and gaseous form. At this temperature, after getting the heat equal to the latent heat of vaporization, water starts changing from liquid state to gaseous state.

## Question 3:

For any substance, why does the temperature remain constant during the change of state?

## Answer 3:

During a change of state, the temperature remains constant. This is because all the heat supplied to increase the temperature is utilised (as latent heat) in changing the state by overcoming the forces of attraction between the particles. Therefore, this heat does not contribute in increasing the temperature of the substance.

## Question 4:

Suggest a method to liquefy atmospheric gases.

## Answer 4:

By applying pressure and reducing the temperature, atmospheric gases can be liquefied.

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## Question 1:

Why does a desert cooler cool better on a hot dry day?

## Answer 1:

When a liquid evaporates, the particles of the liquid absorb energy from the surroundings to compensate the loss of energy during evaporation. This makes the surroundings cool.

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**L Reisser**



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