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web 1. molarity  $M = \frac{\text{mass}}{\text{molar mass} \times \text{volume}}$  4500 g / (825 g/mol  $\times$  0.0129 L) = 2 molality  $m = \frac{\text{mass}}{\text{molar mass} \times \text{volume}}$  825 g / (141.9579 g/mol  $\times$  0.00581158 L) = 0.4500 kg / (0.0129 L  $\times$  3 mole fraction  $n_2 = \frac{\text{mass}}{\text{molar mass}}$  40.825 g / 141.9579 g/mol = 0.2876 mol  $n_1 = \frac{\text{mass}}{\text{molar mass}}$  450.0 g / 18.015 g/mol = 24.97918401 mol

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web 1 calculate the molarity of a solution which contains 0.40 mol of a substance dissolved in 1.6 L of a solution 2 what is the molarity of a solution containing 325 g of NaCl dissolved in 750 ml of solution 3 140 g of KCl is dissolved in 600 ml of water what is the molarity 4 724.4 g of ammonium phosphate in 4500 ml of alcohol

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