

Boyle's Law

1. The volume of a fixed mass of gas varies inversely with the pressure of constant temperature.
2. Boyle's
3. Boyle's Law

Boyle's Theory and Boyle's Law

1. Boyle's of gases is caused by the elasticity of the gas molecules themselves that collide.
2. If the gas is compressible to small the volume is small which increases the pressure which is small.
3. Force exerts pressure on closed on the walls of the container.
4. Pressure of the gas will increase.

Ex. 1. A certain quantity of gas has a volume of 800 ml. at standard atmospheric pressure. After the volume is reduced, it maintains a volume of 400 ml. above the pressure is 1000 mm. What is the volume of the gas at this pressure?

$$P_1 V_1 = P_2 V_2 \quad 1000 \text{ mm} \times 400 \text{ ml} = P_2 \times 800 \text{ ml} = P_2$$

$$P_2 = \frac{1000 \times 400}{800} = 500 \text{ mm}$$

Ex. 2. A gas is pressure of 100 mm. and 400 ml. of gas has a volume of 800 ml. If the pressure is increased until the volume is 200 ml. what is the new pressure, assuming the temperature remains constant?

$$P_1 V_1 = P_2 V_2 \quad 100 \text{ mm} \times 400 \text{ ml} = P_2 \times 200 \text{ ml} = P_2$$

$$P_2 = \frac{100 \times 400}{200} = 200 \text{ mm}$$

Note Taking Guide Episode 902 Chemistry

Katrin Zwirglmaier



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