

## NUCLEAR CHEMISTRY STUDY GUIDE—CH. 21

1. Write the symbol for uranium—238 in isotope format (with top and bottom #s).
2. What does the top # represent?
3. What does the bottom # represent?
4. What particles are found in the nucleus?
5. Why does the strong nuclear force exist?
6. What is the electrostatic force?
7. What particles does the electrostatic force deal with?
8. Define nuclear reactions.
9. Define nuclear fission and fusion.
10. Nuclear reactions usually produce \_\_\_\_\_ amounts of energy, and isotopes become \_\_\_\_\_ stable. The \_\_\_\_\_ number and the \_\_\_\_\_ number may also change.
11. What is transmutation?
12. Write the beta decay of radium—226.
13. Write the alpha decay of radium—226.
14. Why would the isotopes and the end of problems #13 and #14 still be unstable?
15. Define nuclear radiation.
16. List the 3 radioactive particles in order of INCREASING penetrating power.
17. What is radioactive decay?
18. Alpha particles are basically \_\_\_\_\_ nuclei, and beta particles are basically high energy \_\_\_\_\_.
19. What is half life?
20. What is radioactive dating?
21. What is cobalt—60 used for?
22. Does fission or fusion produce nuclei with heavier mass than the original mass?
23. Name 3 differences between chemical and nuclear reactions.
24. What is carbon—13 radioactive?
25. Elements past atomic # \_\_\_\_\_ are radioactive.
26. If  $1/16$  of a sample of a radioactive isotope is left after nuclear decay, how long did this take if the  $\frac{1}{2}$  life is 4 days?
27. If the half life of a radioactive isotope is 4 days, how much of a 50 gram sample is left after 12 days?

# Nuclear Chemistry Study Guide

**RD Boyd**



## **Nuclear Chemistry Study Guide:**

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## **Table of Contents Nuclear Chemistry Study Guide**

1. Understanding the eBook Nuclear Chemistry Study Guide
  - The Rise of Digital Reading Nuclear Chemistry Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Nuclear Chemistry Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Nuclear Chemistry Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Nuclear Chemistry Study Guide
  - Personalized Recommendations
  - Nuclear Chemistry Study Guide User Reviews and Ratings
  - Nuclear Chemistry Study Guide and Bestseller Lists
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  - Nuclear Chemistry Study Guide Budget-Friendly Options
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    - Nuclear Chemistry Study Guide Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Nuclear Chemistry Study Guide
    - Highlighting and Note-Taking Nuclear Chemistry Study Guide
    - Interactive Elements Nuclear Chemistry Study Guide
  8. Staying Engaged with Nuclear Chemistry Study Guide
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Nuclear Chemistry Study Guide
  9. Balancing eBooks and Physical Books Nuclear Chemistry Study Guide
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Nuclear Chemistry Study Guide
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Nuclear Chemistry Study Guide
    - Setting Reading Goals Nuclear Chemistry Study Guide
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Nuclear Chemistry Study Guide
    - Fact-Checking eBook Content of Nuclear Chemistry Study Guide
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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