

NUCLEAR CHEMISTRY WORKSHEET

Name _____ Date _____ Block _____

Most nuclear reactions involve changes in the number of protons and/or neutrons in the nucleus of an atom. These changes are called transmutations because an atom of one element is changed into an atom of a different element. During these changes, nuclear particles may be absorbed and/or emitted along with a release of energy.

There are four types of nuclear reactions:

- 1) natural radioactive decay: transmutations occur by natural radioactivity (the ability of a nucleus to emit a nuclear particle and energy without external stimulation);
- 2) artificial transmutation: transmutations occur by nuclear disintegration caused by external stimulation as a scientist bombards a nucleus with a particle (the addition of another nuclear particle makes the nucleus unstable);
- 3) fission: certain nuclei having a large mass are bombarded with special particles that cause the nuclei to split into two nuclei each having a smaller mass;
- 4) fusion: nuclei of light elements are combined to form heavier nuclei.

Nuclear reactions are written using symbols in the A_ZX notation. The symbols for the major nuclear particles involved in nuclear reactions are given below.

Alpha particle ${}^4_2\text{He}$ (an alpha particle is a helium nucleus)

Proton ${}^1_1\text{H}$ (the most common hydrogen nucleus is a proton)

Neutron ${}^1_0\text{n}$

Electron ${}^0_{-1}\text{e}$ or ${}^0_{+1}\beta$ (also called a beta particle)

Positron ${}^0_{+1}\text{e}$ or ${}^0_{-1}\beta$

Gamma ray ${}^0_0\gamma$

The table below contains information about nuclear reactions involving the emission of radiation.

NUCLEAR CHANGES

Type of Reaction	Radiation (Particle)	Effect on the Atomic Number	Effect on the Atomic Mass
Alpha emission (α)	${}^4_2\text{He}$	decrease by 2	decrease by 4
Beta emission (β^-)	${}^0_{-1}\text{e}$ or ${}^0_{+1}\beta$	increase by 1	no change
Positron emission (β^+)	${}^0_{+1}\text{e}$ or ${}^0_{-1}\beta$	decrease by 1	no change
Gamma emission	${}^0_0\gamma$	no change	no change
Electron capture (EC)	${}^0_{-1}\text{e}$ or ${}^0_{+1}\beta$	decrease by 1	no change

Nuclear Chemistry Answers

John Moore, Richard H. Langley

Nuclear Chemistry Answers:

An Introduction To Nuclear Chemistry M. Satake,2010 Contents The Development of Nuclear Chemistry Fundamental Particles and Nuclear Structure Radioactivity and Nuclear Reactions Properties of Nuclear Radiations The Detection and Measurement of Nuclear Radiation Nuclear Instrumentation Radiation Chemistry Isotope Measurement and Separation Methods Charged Particle Accelerators Neutron Sources Production and the Actinides Uses of Isotopes Experimental Nuclear Chemistry **The Handy Chemistry Answer Book** Justin P. Lomont,Ian C. Stewart,2013-10-01 Don t be mixed up about chemistry Simplify the complex chemical reactions that take place everywhere in our lives with this engaging easy to follow question and answer guide Where would we be without atoms and compounds Gas liquids solids and plasma Acids and bases Bonds and reactions Matter and energy The Handy Chemistry Answer Book covers the building blocks of life and the universe The secret life of atoms how polar bears aren t actually white why oil and water don t mix and much much more are revealed and explained This informative guide covers the basics of chemistry history atomic structures chemical bonds and reactions organic and inorganic chemistry to more advanced material nuclear chemistry biochemistry physical and theoretical chemistry by answering nearly 1 000 common chemistry questions including What causes lightning How does photosynthesis work What are hard and soft Lewis acids and bases What makes a fabric waterproof What are the twelve principles of green chemistry When did alchemists finally abandon trying to make gold What is Le Chatelier s principle What do the different octane ratings mean at the gas pump What is genetic engineering Why is calcium important for strong bones What is the 18 electron rule Why does chocolate turn white as it ages Chemical reactions that rule the world their properties structure composition behavior and history are tackled and explained in plain English in The Handy Chemistry Answer Book With many photos illustrations a few formulas molecular diagrams and other graphics this fun fact filled tome is richly illustrated A history of chemistry timeline appendices on Nobel Prize in Chemistry winners a bibliography further reading section glossary of terms a table of physical constants a table of conversion factors and extensive index add to its usefulness *Objective Question Bank in Chemistry* B.K. Sharma,1997 *E3 Chemistry Review Book - 2018 Home Edition (Answer Key Included)* Effiong Eyo,2017-10-20 With Answer Key to All Questions Chemistry students and homeschoolers Go beyond just passing Enhance your understanding of chemistry and get higher marks on homework quizzes tests and the regents exam with E3 Chemistry Review Book 2018 With E3 Chemistry Review Book students will get clean clear engaging exciting and easy to understand high school chemistry concepts with emphasis on New York State Regents Chemistry the Physical Setting Easy to read format to help students easily remember key and must know chemistry materials Several example problems with solutions to study and follow Several practice multiple choice and short answer questions at the end of each lesson to test understanding of the materials 12 topics of Regents question sets and 3 most recent Regents exams to practice and prep for any Regents Exam This is the Home Edition of the book Also available in School Edition ISBN

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McGraw-Hill's SAT Subject Test: Chemistry, 2ed Thomas Evangelist,2009-02-01

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Segal,Peter S. Shenkin,1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry Worked examples now number over 300 and exercises over 1460

Principles Of Nuclear

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