

Practice Problems Unit 5 – Nuclear Chemistry

Directions: Answer each of the following questions. Show ALL work in the space provided with units and correct sig figs to receive credit.

Practice Problems – Nuclear Reactions

1. Fill out the following chart to help predict types of radiation.

Type of Radiation	How can you recognize that problem?		
beta decay	${}_{7}^{14}\text{Ne}$	mass # 14	atomic # 17
alpha decay	${}_{8}^{14}\text{He}$	mass # same	atomic # 21
gamma decay		γ - no change	

2. Fill out the chart of based on your knowledge of nuclear reactions.

Element	Protons	Neutrons	Electrons	Stable?	Type of Radiation	Balanced Nuclear Reaction
carbon-14	6	8	6	no	beta decay	${}_{6}^{14}\text{C} \rightarrow {}_{7}^{14}\text{N} + {}_{-1}^0\beta$
radium-226	88	$226 - 88 = 138$	88	no	alpha decay	${}_{88}^{226}\text{Ra} \rightarrow {}_{2}^{4}\text{He} + {}_{86}^{222}\text{Rn}$
${}^{65}\text{Cu}$	29	$65 - 29 = 36$	29	no	beta decay	${}_{29}^{65}\text{Cu} \rightarrow {}_{-1}^0\beta + {}_{30}^{65}\text{Zn}$
${}^{127}\text{I}$	53	$127 - 53 = 74$	53	no	gamma decay	${}_{53}^{127}\text{I} \rightarrow {}_{-1}^0\gamma + {}_{53}^{127}\text{I}$
${}^{186}\text{Re}$	75	$186 - 75 = 111$	75	no	gamma decay	${}_{75}^{186}\text{Re} \rightarrow {}_{-1}^0\gamma + {}_{75}^{186}\text{Ru}$
nobelium-259	102	$259 - 102 = 157$	102	no	alpha decay	${}_{102}^{259}\text{No} \rightarrow {}_{2}^{4}\text{He} + {}_{100}^{255}\text{Tm}$

3. How are the mass number and atomic number of the nucleus affected by the following:

a. Beta decay mass # same atomic # 11

b. Alpha decay mass # 1/4 atomic # 2

c. Gamma decay mass # same atomic # same

4. Complete and balance each of the following nuclear reactions.



5. Bromine -82 decays to produce Krypton - 82.

a. What type of decay is this? Justify your answer! beta decay - mass stayed the same + atomic # 11

b. Write the nuclear reaction for this decay. ${}_{35}^{82}\text{Br} \rightarrow {}_{36}^{82}\text{Kr} + {}_{-1}^0\beta$

6. Neptunium - 237 undergoes decay to Protactinium - 233.

a. What type of decay is this? Justify your answer! alpha decay \rightarrow mass # 1/4 + atomic # 2

b. Write the nuclear reaction for this decay. ${}_{93}^{237}\text{Np} \rightarrow {}_{91}^{233}\text{Pa} + {}_{2}^{4}\text{He}$

7. Write a balanced nuclear reaction for the beta decay of:



8. Write a balanced nuclear reaction for the alpha decay of:



Nuclear Chemistry Test Answers

Darrell D. Ebbing

Nuclear Chemistry Test Answers:

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