

# Practice Problems Unit 5 – Nuclear Chemistry

**Directions:** Answer each of the following questions. Show ALL work in the space provided with units and correct sig figs to receive credit.

## Practice Problems – Nuclear Reactions

1. Fill out the following chart to help predict types of radiation.

Type of Radiation	How can you recognize that problem?		
beta decay	$\beta$ or $e^-$	mass # $\downarrow 4$	atomic # $\downarrow 1$
alpha decay	$\alpha$ or $He$	mass # $\downarrow 4$	atomic # $\downarrow 2$
gamma decay	$\gamma$	no change	

2. Fill out the chart based on your knowledge of nuclear reactions.

Element	Protons	Neutrons	Electrons	Stable?	Type of Radiation	Balanced Nuclear Reaction
carbon-14	6	8	6	no	beta decay	$^{14}_6C \rightarrow ^{14}_7N + ^0_{-1}\beta$
radium-226	88	$226 - 88 = 138$	88	no	alpha decay	$^{226}_{88}Ra \rightarrow ^4_2He + ^{222}_{86}Rn$
$^{65}_{29}Cu$	29	$65 - 29 = 36$	29	no	beta decay	$^{65}_{29}Cu \rightarrow ^{65}_{30}Zn + ^0_{-1}e$
$^{127}_{53}I$	53	$127 - 53 = 74$	53	no	gamma decay	$^{127}_{53}I \rightarrow ^{127}_{53}I + \gamma$
$^{186}_{75}Re$	75	$186 - 75 = 111$	75	no	gamma decay	$^{186}_{75}Re \rightarrow ^{186}_{75}Re + \gamma$
nobelium-259	102	$259 - 102 = 157$	102	no	alpha decay	$^{259}_{102}No \rightarrow ^4_2He + ^{255}_{100}Fm$

3. How are the mass number and atomic number of the nucleus affected by the following:

- Beta decay mass # same atomic #  $\uparrow 1$
- Alpha decay mass #  $\downarrow 4$  atomic #  $\downarrow 2$
- Gamma decay mass # same atomic # same

4. Complete and balance each of the following nuclear reactions.

- $^{28}_{13}Al \rightarrow ^{28}_{14}Si + ^0_{-1}e$
- $^{210}_{83}Bi \rightarrow ^{206}_{81}Tl + ^4_2He$

5. Bromine -82 decays to produce Krypton -82.

- What type of decay is this? justify your answer! beta decay - mass stayed the same + atomic #  $\uparrow 1$
- Write the nuclear reaction for this decay.  $^{82}_{35}Br \rightarrow ^{82}_{36}Kr + ^0_{-1}e$

6. Neptunium - 237 undergoes decay to Protactinium - 233.

- What type of decay is this? justify your answer! alpha decay  $\rightarrow$  mass #  $\downarrow 4$  + atomic #  $\downarrow 2$
- Write the nuclear reaction for this decay.  $^{237}_{93}Np \rightarrow ^{233}_{91}Pa + ^4_2He$

7. Write a balanced nuclear reaction for the (beta decay) of:

- Nitrogen - 16  $^{16}_7N \rightarrow ^{16}_8O + ^0_{-1}e$
- Potassium - 40  $^{40}_{19}K \rightarrow ^{40}_{20}Ca + ^0_{-1}e$

8. Write a balanced nuclear reaction for the (alpha decay) of:

- Plutonium - 244  $^{244}_{94}Pu \rightarrow ^4_2He + ^{240}_{92}U$
- Strontium - 90  $^{90}_{38}Sr \rightarrow ^4_2He + ^{86}_{36}Kr$

# Nuclear Chemistry Test Answers

**Darrell D. Ebbing**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## Nuclear Chemistry Test Answers:

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## **Table of Contents Nuclear Chemistry Test Answers**

1. Understanding the eBook Nuclear Chemistry Test Answers
  - The Rise of Digital Reading Nuclear Chemistry Test Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Nuclear Chemistry Test Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Nuclear Chemistry Test Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Nuclear Chemistry Test Answers
  - Personalized Recommendations
  - Nuclear Chemistry Test Answers User Reviews and Ratings
  - Nuclear Chemistry Test Answers and Bestseller Lists
5. Accessing Nuclear Chemistry Test Answers Free and Paid eBooks
  - Nuclear Chemistry Test Answers Public Domain eBooks
  - Nuclear Chemistry Test Answers eBook Subscription Services



- Nuclear Chemistry Test Answers Budget-Friendly Options
- 6. Navigating Nuclear Chemistry Test Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Nuclear Chemistry Test Answers Compatibility with Devices
  - Nuclear Chemistry Test Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Nuclear Chemistry Test Answers
  - Highlighting and Note-Taking Nuclear Chemistry Test Answers
  - Interactive Elements Nuclear Chemistry Test Answers
- 8. Staying Engaged with Nuclear Chemistry Test Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Nuclear Chemistry Test Answers
- 9. Balancing eBooks and Physical Books Nuclear Chemistry Test Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Nuclear Chemistry Test Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Nuclear Chemistry Test Answers
  - Setting Reading Goals Nuclear Chemistry Test Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Nuclear Chemistry Test Answers
  - Fact-Checking eBook Content of Nuclear Chemistry Test Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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